OSHIBKOV, V.

Films on progressive practice are needed. Muk.-elev.prom. 20 no.3:31 Mr '54. (MLRA 7:7)

1. Tambovskoye upravleniye BOAMVe.
(Grain--Storage)

OSHIBKOV, V.

Instructions for the disinfection of peas need to be revised. Muk.-elev.prom. 21 no.5:30 My '55. (MLRA 8:9)

1. Tambovskaya kontora Zagotzerno. (Peas--Disinfection)

OSHIGOV, T. P.

"Distribution of the [illegible] periodic system." (p. 1747)

SO: Journal of General Chemistry (Zhurnal obshchei khimii), 1951, [illegible]

Oshigov, E. P.

ZHURNAL ANALITICHESKOY KHIMII

Vol 11, No. 3, May-June, 1956

ON LEAD DETECTION IN ORES BY THE GRINDING METHOD

E. P. Oshigov, M. A. Rafienko and *L. K. Ivanenko*

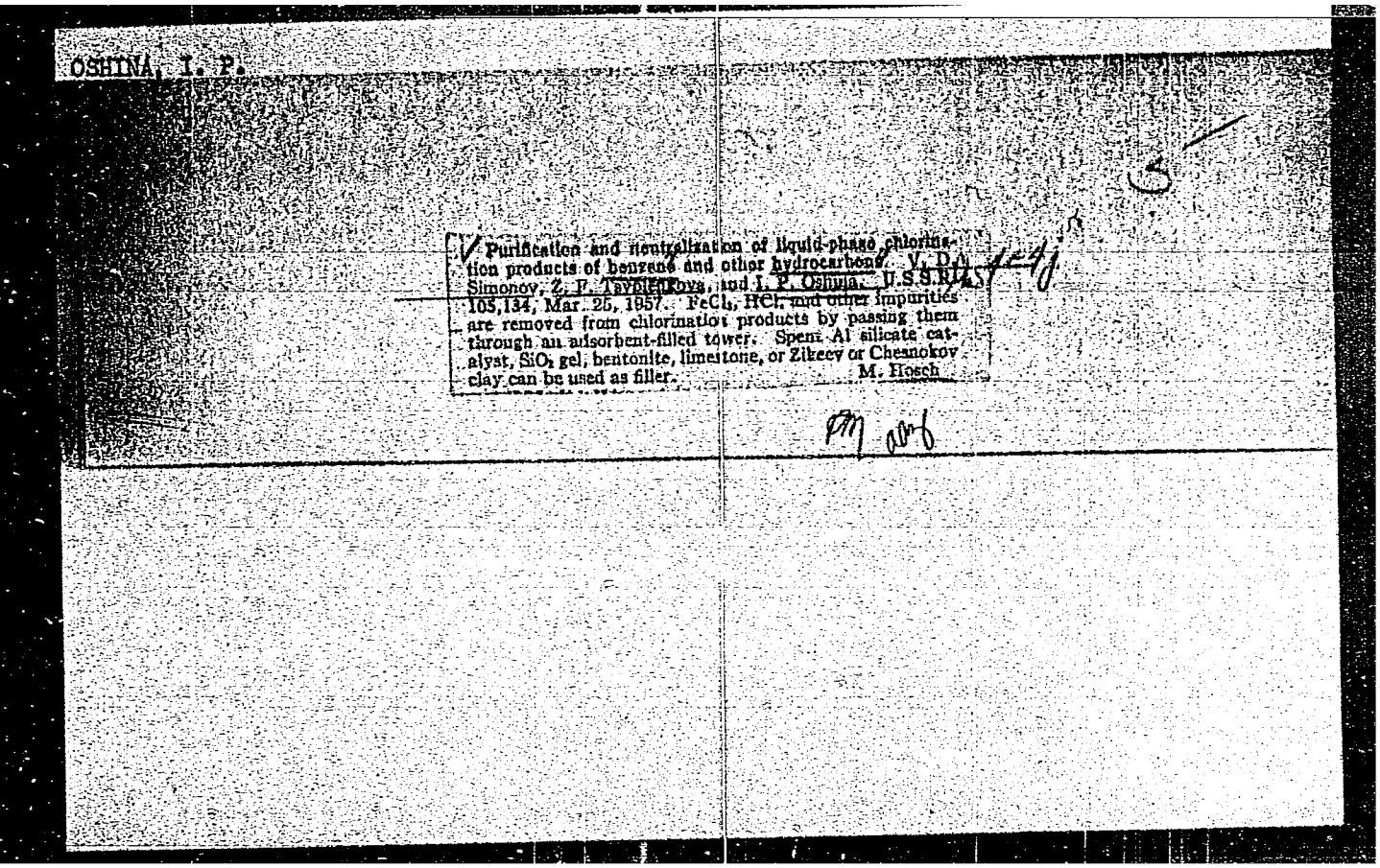
Far-Eastern Branch of the Academy of Sciences USSR bearing the name of V. L. Komarov

Lead detection by means of K[using the grinding method with preheating in the last stage shows high specifity and sensitivity. The reaction may be used for lead detection not only in lead minorals and concentrates but also for the analysis of complex base metallic and fluorite ores containing a great number of various elements.

SKOROKHODOV, N.Ye., kandidat tekhnicheskikh nauk, dotsent; GOLUBEV, T.M., professor, doktor tekhnicheskikh nauk; ZAYKOV, M.A., kandidat tekhnicheskikh nauk; CHELYSHEV, N.A., kandidat tekhnicheskikh nauk, dotsent; KOROLEV, A.S., inzhener; OSHIN, V.I., inzhener.

Determining acting forces in friction and eccentric presses. Trudy Sib.met.inst. no.2:19-29 '55. (MLRA 9:12)

(Strains and stresses) (Power presses)



OSHINA, I. P.

Purification and neutralization of liquid-phase chlorination products of benzene and other hydrocarbons. V. D. Simonov, Z. F. Tsvetkova, and I. P. Oshina. U.S.S.R. 105,134, Mar. 25, 1957. FeCl$_3$, HCl, and other impurities are removed from chlorination products by passing them through an adsorbent-filled tower. Spent Al silicate catalyst, SiO$_2$ gel, bentonite, limestone, or Zikeev or Chesnokov clay can be used as filler. M. Hosch

[illegible], [illegible]

[illegible]: "[illegible] [illegible] ... [illegible]cial [illegible]-[illegible]tion". [illegible], [illegible]. Latvian [illegible], [illegible]. [illegible] for the Degree of Cand[illegible]

[illegible]: [illegible] [illegible]', [illegible], [illegible] 1[illegible]. [illegible].

OSHIS, F.

Conditions of formation of various modifications of calcium carbonate. A. Ievinš, P. Oša, and B. Mateuss., *Latvijas PSR Zinātņu Akad. Vēstis* 1953, No. 2 (Whole No. 91), 87–100 (in Russian).—$CaCO_3$ was ppt'd. from satd. $Ca(OH)_2$ soln. by bubbling CO_2 until pH 8.2 was reached. Between 10 and 70°, the ppt. was calcite (I), with a trace of aragonite (II) towards higher temp. Crystals were rhombohedric, 3–15 μ at 10°; prismatic at higher temp., with length 7–11 μ at 50°, and becoming shorter with increase in temp. At 80–90°, the ppt. was mostly II, in forms of 2–4-μ platelets and prisms. Tech. CaO and chemically pure $Ca(OH)_2$ gave similar results. In the presence of sulfates of Ca, Na, Mg, and NH_4, II was formed from 65° up; at 10–30°, the crystals were oval or grainlike, 0.5–6 μ; at 60°, mainly prisms with pointed ends; at higher temp., needles up to 22 μ long. In the presence of Cl^-, the ppt. was 0.3–0.7 μ rhombohedral crystals with high percentage of II. In similar formation of $CaCO_3$ from 5–10% suspensions of $Ca(OH)_2$, 0.3–2-μ prisms or platelets of I were formed at 35–80° regardless of the presence or absence of $CaSO_4$, $(NH_4)_2SO_4$, or NaCl. Only a trace of II was observed at 70–80°. It is concluded that $CaCO_3$ is formed by reaction of Ca^{++} with CO_3^{--} in $Ca(OH)_2$ solns., but by a reaction of solid $Ca(OH)_2$ with CO_2 in the case of suspension. The specific surface of the ppt., measured by the air-permeation velocity, was 3700–10200 sq. cm./g.; it increased with the pptn. temp. and decreased with concn. of SO_4^{--}. A. P.

(2)

15-57-12-17345

Translation from: Referativnyy zhurnal, Geologiya, 1957, Nr 12, pp 97-98 (USSR)

AUTHORS: Oshis, F., Iyevinysh, A.

TITLE: Varieties of Calcium Carbonate (Razlichnyye modifikatsii karbonata kal'tsiya)

PERIODICAL: Uch. zap. Latv. un-t, 1956, Vol 9, pp 3-21

ABSTRACT: Calcium carbonate (KK) occurs naturally in the form of calcite and aragonite; aside from these, the following less stable anhydrous crystalline modifications of KK are produced experimentally: waterite of the trigonal system and mi-KK of the hexagonal system. Both of these are sometimes found in nature. At the temperatures below 20°, there exists a monoclinic hexahydrate of KK and an amorphous KK in the form of a gel. An experimental study has been conducted to determine conditions under which these varieties are formed. Carbonization by means of carbon monoxide was tried out, with the gas obtained from limestone or from a chemically pure

Card 1/4

Varieties of Calcium Carbonate (Cont.)

calcium hydroxide. It was established that the variety of calcium carbonate produced, its habit, size, type of aggregation and specific surface depended on the temperature of carbonization and on the presence of extraneous matter. If the carbonizing atmosphere contained admixtures which could serve as nuclei of crystallization, then calcite was formed at all the temperatures up to 100°; in the absence of such nuclei, amorphous KK was precipitated from the solution. The purest waterite was obtained at the temperature of 20° in the absence of nuclei of crystals, if the carbonization was discontinued at the moment when the atmosphere became neutral. With the rise of temperature the content of rhombohedral calcite crystals was increased and waterite decreased. At 40°, somatic forms appear, at over 40° prisms are developed on crystals of calcite instead of rhombohedrons, which fact is augmented by an increase in the duration of carbonization. Calcite--the most stable modification--is also formed out of less stable modifications: at below 20° calcite is formed from hexahydrite, above this temperature calcite can be formed either from supersaturated solutions or by a crystallization of gel of KK. Aragonite is formed at increased

temperatures. KK, containing a small amount of aragonite, is formed during a carbonization of calcium hydroxide at 70°; at 80° to 90°, the amount of aragonite increases. Aragonite (free of calcite) is successfully obtained when a solution of calcium bicarbonate (which is saturated at a normal temperature) is heated so that the temperature of crystallization is near 100°. When a pure aqueous solution of calcium bicarbonate is heated, mi-carbonate in the form of small hexagonal stars and lamellae with Np of 1.550 is obtained at 60° to 65°. At this temperature, under acid conditions, the product is comparatively stable. Waterite obtained in the form of spherolites by carbonizing a solution of calcium hydroxide appears to be uniaxial and negative; its Ng is 1.584. Waterite and mi-KK produce identical X-ray pictures. At 419° to 485° waterite passes into calcite. Its thermal curve shows an exothermic maximum at 459°. Transition of the Ural aragonite into calcite takes place at 444° to 504° (endothermic minimum at 472°). The presence of soluble chlorides and sulfates affects the form of crystals and changes their surface. Chlorides increase the number of crystals and decrease their sizes. Sulfates cause the formation of somatic forms similar

Varieties of Calcium Carbonate (Cont.)

to aragonite. Albumins and carbohydrates increase the solubility of calcium hydroxide and at an increased temperature aid the formation of stable modifications of the investigated compounds. The presence of Mg ion neither interferes nor aids in the formation of aragonite.

Card 4/4 Ye. Ye. Kostyleva

68274

5.1190

Translation from: Referativnyy Zhurnal. Khimiya, 1959, Nr 10, p 295 (USSR)

AUTHORS: Osipov, L., Oshis, F., Kimene, I.

TITLE: Catalytic Decomposition of Ammonia

PERIODICAL: Uch. zap. Latv. un-t, 1958, Vol 22, pp 101-106

ABSTRACT: The thermal decomposition of NH_3 has been investigated under the dynamic conditions of the gaseous phase in the presence of catalysts, reduced natural magnetite, Fe-chips, ferrosilicon, activated ferrosilicon, synthetic catalyst containing the oxides of K, Fe and Al, etc. The most active catalyst proved to be magnetite, in the presence of which NH_3 is decomposed in quantities already at 600°C. There are 17 references

From the authors' summary

Card 1/1

0515, 2

Colloidchemical phenomena at surfaces of metals and retardation of corrosion in salt solutions. I. Distribution of aluminum corrosion in potassium salt solutions. J. Liepiņš and Z. Osis (Chem. Inst., Acad. Sci. Latv. S.S.R., Riga). *Latvijas Zinātņu Akad. Vēstis* 1952, No. 8 (Whole No. [illegible]), 107–13(in Russian; in Latvian, 114–16); cf. *C.A.* 48, [illegible], 9300f.—Vertically suspended Al samples were allowed to corrode in 0.001–2*N* KCl, and the distribution of corrosion forms was studied microscopically. Four forms of corrosion products were observed: small hills or craters (I), glassy spherical accumulations (II), loose accumulations resembling waves (III), and thin adherent layers (IV). The no. of I increased with KCl concn., reaching a max. at 0.5–1*N*, and then decreased. Similar variation was observed in pptn. of Cu on Al from $CuSO_4$ solns. The formation of the craterlike elevations is explained: the highly polymerized forms of Al hydroxides are stabilized at the anodic region with Al^{+++} and similar positively charged ions; the positively charged particles move electrophoretically to the cathodes on the surface of the initial oxide film, meet OH^- there, and are partly dissolved with formation of aluminates and partly recharged to become neg. by ad-

compositions of the products of iron corrosion in solutions of chlorides of alkali and alkaline earth metals. *Ibid.* 1953, No. 11(Whole No. 76), 119–29(in Russian; Latvian summary, 130).—Steel was corroded in 0.001*N*–2*N* solns. of NaCl, KCl, $MgCl_2$, and $CaCl_2$ for several months, and the corrosion products (I) were studied by x-rays. The amt. of the insol. I was relatively independent of the nature of the cation, but changed with concn. reaching a max. at 0.001*N*. The pH was 6.0–6.5 at the beginning, and 6.5–6.8 at the end of the expt. The color of I was yellow-brown in dil. solns. and red-brown in concd. solns. I were highly dispersed, especially at high salt concn., and one dark and one light constituent could be distinguished microscopically. The basic insol. corrosion product in all solns. and in pure water was γ-FeO(OH) (II), identical with the mineral lepidocrocite. The second solid phase was magnetite, $Fe^{II}(Fe^{III}O_2)_2$ (III), formed by a reaction of II, or of the sol. Fe ion, with it. With an increase in concn. of Cl^-, the relative amts. of III decreased. The amts. of III varied with the type of cation as follows: Mg > Na,K > Ca. The concn. of sol. Fe ions was between 1.5×10^{-4} to 1×10^{-6}*M*, increasing with the salt concn. It is proposed that the higher dispersity of the [illegible] charged (by Fe ions) colloidal corrosion products

transferring into bayerite and by[illegible] ...son for the retardation of corrosion in the concrete [illegible]
little effect on the rate of corrosion; most retardation is caused by the formation of I. VII. Chemical and phase
Andrew Draviņš

OSIS, Z.F.

1335. Osis, Z. F. Khim'cheskiy i fazovyy sostav produktov korrozii alyuminiya i zheleza v rastvorakhsole. Riga, 19[illegible]. 1 - 22 s. (latv. gos. un-t. khim. fak.) 150 zkz. B. ts. ----(Ab---9[illegible]

SO: Knizhnaya Letopis, Vol. 1, 19[illegible]

0515; Z

Colloid-chemical phenomena on the surfaces of metals and retardation of corrosion in salt solutions. VIII. Corrosion of aluminum and iron as a function of the concentration of solution. L. Liepiņa, A. Vaivade, Z. Oša, and A. Stiprais (Inst. Chem., Acad. Sci. Latv. SSR, Riga, Latvia). *Latvijas PSR Zinātņu Akad. Vestis* 1954, No. 3 (Whole No. 80), 107–13 (in Russian); cf. *C.A.* 47, 11921b. —In 0.001–3*N* KCl at 20°, the long-range (50 days) corrosion rate of Al follows the equation $\Delta g = kc^{\alpha}$, where Δg is the wt. loss, c concn. of KCl, and k and α are consts., e.g. 0.8 and 0.43, resp. Similar relations hold for Al and K sulfate solns., but here the corrosion is slower. In shorter exposures, e.g. 15 days in KCl, this law is not obeyed and a max. corrosion rate is observed at 1*N* concn. An explanation for the equation is proposed: at higher concns. of Cl^-, the polyoxychlorides, formed probably through the intermediate steps of adsorption on the primary hydroxide (boehmite) and ion exchange, peptize the primary hydroxide deposits; in dil. solns., the primary boehmite passes into the less reaction-active bayerite and hydrargillite which block the surfaces and slow down the corrosion. Resemblance between the above equation and relations expected from the adsorption isotherm and mass-action law are pointed out. In the corrosion of Fe, the short-time corrosion is higher in the more dil. solns.; chlorides and sulfates of the alkali and alkaline earth metals were investigated. In long-time tests, the corrosion rate was max. in 0.1–0.001*N* solns. This is explained by blocking of the cathodic areas by an electrophoretic pptn. of the positively charged colloidal particles of the hydroxidic primary and secondary corrosion products. The irregularities are explained by changes in the rates of formation and aging of the secondary corrosion products such as magnetite. The relative amts. of the latter in the corrosion products increased with the dilu. of the electrolyte. The magnetite forms by a reaction of the bivalent Fe ions with the secondarily formed (by oxidation) trivalent hydroxide lepidocrocite. Andrew Dravnieks

USSR

Colloid-chemical phenomena at surfaces of metals and retardation of corrosion in salt solutions. X. Chemical and phase composition of products of corrosion of lead in water and in solutions of potassium chloride. Z. Ošis and L. Liepiņa. *Latvijas PSR Zinātņu Akad. Vēstis* 1954, No. 12(Whole No. 89), 125–32(in Russian; Latvian summary); cf. *C.A.* 49, 5350g.—The product of corrosion of Pb in distd. H_2O exposed to air was $2PbCO_3.Pb(OH)_2$ (I), and the pH of the water during the corrosion process was stabilized at 8.0. In solns. of KCl below 0.001*N*, the corrosion products were also I. In 0.001*N* to 0.5*N* solns., the products were I and an oxychloride, tentatively identified as $2Pb(OH)_2.PbCl_2$ (II), but definitely not Pb(OH)Cl. Above 0.5*N*, the products were I and H_2O-sol. complex Pb chlorides (III). Below 0.5*N*, $PbCO_3$, resulting from interaction with CO_2 from air, was also detected in the products. The max. in the corrosion rate vs. KCl concn. curve at 0.05*N* soln. is explained by formation of coarse cryst. II, and another max. at concn. above 2*N* by formation of III. In 0.5*N* to 2.0*N* $KHCO_3$, the corrosion product was $PbCO_3$. XII. Corrosion of steel and its retardation in solutions of neutral and acid carbonates ($KHCO_3$ and K_2CO_3). L. Liepiņa and A. Valvads. *Ibid.* 1955, No. 1(Whole No. 90), 131–7(in Russian; Latvian summary); cf. *C.A.* 49, 6071*b*.—Corrosion of common steel was studied under static conditions in solns. of K_2CO_3 (I) and $KHCO_3$ (II), open to air at 20°. In solns. of I, the corrosion rate (4.8 mg./sq. cm./month) did not vary with concn. of I from 0 to 0.6*N*; the

1/2

3; (Aug.)

corrosion was linear with time, and the corrosion products were yellow-brown and did not adhere to the surface of steel. Above 0.025*N*, the products became greenish brown, adhered to steel, and the corrosion became slower with increase of concn. of I; above 0.1*N* I no corrosion was observed and the pH of the soln. was above 10. In solns. of II, corrosion products became more amorphous and more adherent at concn. above 0.1*N*, and corrosion was inhibited in solns. above 0.8*N* at lower pH than in solns. of I. It is proposed that the primary corrosion product in dil. solns. is sol. $Fe(HCO_3)_2$, which is oxidized to insol. α-FeO(OH) (III) at some distance from the surface of the steel, and therefore cannot retard corrosion efficiently. In more concd. solns. and at higher pH, the primary product is insol. $FeCO_3$, and the secondary III may form in the immediate vicinity of the anodic areas. The influence of pH is explained by observations that at pH 11, the stable form is III rather than γ-FeO(OH) (IV). Since the colloid particles of III are positively, but those of IV negatively charged, an electrophoretic blocking of anodic areas is possible with III but not with IV.

Andrew Dravnieks

3/2

OSHIS, Z. F.

OSHIS, Z. F. -- "Chemical and Phase Composition of the Corrosion Products of Aluminum and Iron in Salt Solutions." Latvian State U, 1955 In Latvian (Dissertation for the Degree of Candidate of Chemical Sciences)

SO: Izvestiya Ak. Nauk Latvivskoy SSR, No. 9, Sept., 1955

USSR

9502* Relation of the Corrosion Rate of Iron to the pH of Solution, and the Passivation of the Metal in Alkaline Solutions. **Zavisimost' skorosti korrozii zheleza ot pH rastvora i passivatsiia metalla v shchelochnykh rastvorakh.** (Russian.) L. K. Lepin', A. Ia. Vaivade, and Z. F. Oshis. *Zhurnal Fizicheskoi Khimii*, v. 29, no. 2, Feb. 1955, p. 350-356 + 1 plate.

Oxidation kinetics of Fe in various media; transition from γ to α form. Retardation of corrosion at certain pH values. Graphs, table, diffraction patterns. 10 ref.

...ticles of viranite. A. Dravnieks

SOV/76-33-2-...

AUTHORS: Lepin', L. K., Vagvale, L. Ya., Ozols, Z. F.

TITLE: Oxidation Kinetics of Iron in Aqueous Solutions of the Salt Mixtures [$KCl(K_2SO_4)+K_3PO_4(K_2HPO_4,K_2CO_3)$] (Kinetika okisleniya zheleza v rastvorakh smesey soley [$KCl(K_2SO_4)+K_3PO_4(K_2HPO_4,K_2CO_3)$])

PERIODICAL: Zhurnal fizicheskoy khimii, 1959, Vol 33, Nr 2, pp 357 - 366 (USSR)

ABSTRACT: On the basis of observations in previous papers (Refs 1-5) it can be assumed that the prevention of iron corrosion in neutral salt solutions of alkali metals (chlorides, sulfates) occurs through an electrophoretic blocking of the cathodic segments of the metal surface by the positively charged γ-FeO(OH) particles. This assumption is corroborated by the determinations of the electrode potential of iron (Refs 6,7). The iron oxidation in carbonate and phosphate solutions of the alkali metals is nevertheless hindered by negatively charged αFe(OH)₃ particles and this can finally lead to a passivation of the metal surface (Refs 2-6). For this reason it seemed interesting to carry out investigations

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with mixtures of the above mentioned salts. A few important papers concerning this matter are given, among which are those by I.Shtern, Ts. Gul'yanskaya and K.Nekrasov (Ref 11), V. A. Rozenberg and Ye. I. Logorel'skiy (Ref 13), I. L. Rozenfel'd (Ref 14), and others (Refs 8-10, 12), and it is found that the prevention of corrosion must depend upon the proportional amount of the salt components. For this reason the kinetics and the character of the iron corrosion were investigated for solutions of KCl, K_2SO_4, K_2HPO_4, K_3PO_4 and K_2CO_3, and binary solutions of these salts at 20°C. For these studies steel 10(C - 0.13%, Si-0.28%, Mn-0.55%, P - 0.036%, S - 0.042%, the rest Fe) was used. It was observed that an increase in the phosphate or carbonate concentration (Figs 3-5) with a constant concentration of KCl or K_2SO_4 (under 1.0 n) accelerates the corrosion at the beginning, then this effect passes through a maximum and fades with a ratio of the inhibitor to the salt of 5(10) : 1 to exhibit a passivating effect. At concentrations

Card 2/3

Dissolution Kinetics of Iron in Aqueous Solutions of the Salt Mixtures $[KCl(K_2SO_4)+K_2PO_4(K_2HPO_4,K_2CO_3)]$ SOV/76-33-2-19/49

of KCl > 1.0 n no maximum appears and the corrosion is not completely inhibited. It is assumed that with small salt concentrations an inhibition of the corrosion occurs according to the above assumption, by γ-FeO(OH) particles, while at higher concentrations of the inhibitor negative particles block the anodic segments. The observed maximum on the corrosion-concentration curves is explained in terms of an over-charging of the particles, which block the metal surface. There are 6 figures, 2 tables, and 17 references, 12 of which are Soviet.

ASSOCIATION: Akademiya nauk Latv. SSR, Institut khimii (Academy of Sciences Latv. SSR, Institute for Chemistry)

SUBMITTED: July 17, 1957

Card 3/3

CHEMODANOV, V.S.; OSHITKO, V.M.; SULTANOV, S.A.; VAKHITOV, G.G.; POLUYAN, I.G.

Conversion of reserves and the determination of the recovery factor of a flooded section of reservoir D_1 in the Bavly field. Neftepromo delo no.1:13-15'63 (MIRA 17:7)

1. Tatarskiy neftyanoy nauchno-issledovatel'skiy institut, g. Bugul'ma i Neftepromyslovoye upravleniye "Bavlyneft'".

DEYEV, M.Ya., master; YELCHEV, G.A., slesar'; SNIGIREV, F.I., slesar'; KZKRASOV, V.G., slesar'; NAD'KIN, N.A., mashinist elektrovoza; OSHIVALOV, A.V., mashinist elektrovoza; PANCHENKO, P.M., mashinist elektrovoza.

Brush-holder units must be improved. Elek.i tepl.tiaga 2 no.4:6-7 Ap '58. (MIRA 12:3)

1. Elektromashinnyy tsekh depo Zlatoust Yuzhno-Ural'skoy dorogi (for Deyev). 2. Depo Zlatoust-Yuzhno-Ural'skoy dorogi (for all except Deyev).

(Electric brushes) (Electric railway motors)

MISHKOV, Yu.Ya.; OSHKADEROV, S.P.

Methods of temperature recording during ultra high-speed heating of iron. Sb. nauch. rab. Inst. metallurgii AN URSR no.18[illegible] (MIRA 1[illegible]8)

GRIDNEV, V.N.; MESHKOV, Yu.Ya.; OSHKADEROV, S.P.

Critical points of pure iron during rapid electric heating.
Sbor. nauch. rab. Inst. metallofiz. AN URSR no.18:96-99 '64
(MIRA 17:8)

L 36621-65 EWT(m)/EWA(d)/T/EWP(t)/EWP(b)/EWA(c) MJW/JD
ACCESSION NR: AP5002356 S/0126/64/018/006/0938/0939

AUTHOR: Gridnev, V. N.; Meshkov, Yu. Ya.; Oshkaderov, S. P.

TITLE: Temperature of austenite formation during rapid heating of steel and iron

SOURCE: Fizika metallov i metallovedeniye, v. 18, no. 6, 1964, 938-939

TOPIC TAGS: austenite, austenite formation, formation temperature, iron, U8A steel, nondiffusion transformation

ABSTRACT: The temperatures at which austenite was formed from pure iron, or from tempered, normalized or annealed U8A steel by heating at different rates (20-6000 deg/sec for steel and up to 10000 deg/sec for iron) were determined. The moment of the α -γ transformation was determined by dilatometric analysis which indicated compression changes on an oscillogram. The study con-

ed, and indicated the ... the greater the dependence on the ... the coarser the initial structure,

Card 1/2

L 36621-65
ACCESSION NR: AP5002356

Since the rise of the critical point in steel was much more rapid than in iron, it appeared theoretically possible to coincide the critical points in iron and in pearlite by heating the latter sufficiently rapidly. This was not achieved experimentally. Absolute nondiffusion transformation by heating annealed steel is apparently not possible because in the 723-905C range some portion of the austen-

...ecomplexes) by a recombination mechanism. Orig. art. has: 1 figure and 1 equation.

ASSOCIATION: Institut metallofiziki AN UkrSSR (Institute of Metallophysics, AN UkrSSR)

SUBMITTED: 10Mar64 ENCL: 00 SUB CODE: MM

NR REF SOV: 008 OTHER: 000

Card 2/2

GRIDNEV, V. N.; GORKADIROV, S.P.; TARASOV, V.F.

Method of the differentiation of electric resistance ... electric heating. Zav. lab. 31 no.2 ...

1. Institut metallofiziki AN UkrSSR.

GRIDNEV, V.N.; ME[illegible]HKOV, Yu.Ye.; [illegible]SHRADEROV, S.F.

[illegible]omposition of austenite formed during the rapid heating of steel. Sbor.nau[illegible] trud. Inst. metall fiz. AN URSR no.[illegible] [illegible] 1[illegible].
(MIRA [illegible])

GRIDNEV, V.N.; MESHKOV, Yu.Ya.; OSHKADEROV, S.P.

Diagram of the isothermal formation of austenite in steel.
Sbor. nauch. trud. Inst. metallofiz. AN UkrSSR no.20:142-153
'64. (MIRA 18:5)

MESHKOV, Yu.Ya.; OSHKAYEROV, S.P.; CHERNENKO, V.V.

Equipment for complex investigations of phase transformations under the effect of rapid heating. Sor. nauch. trud. Inst. metallofiz. AN URSR no.26:198-202 '64.

(MIRA 18:5)

SHKUTA, E.I.; LUGOVSKIY, S.I., doktor tekhn.nauk; OSHMYANSKIY, I.B., gornyy inzh.

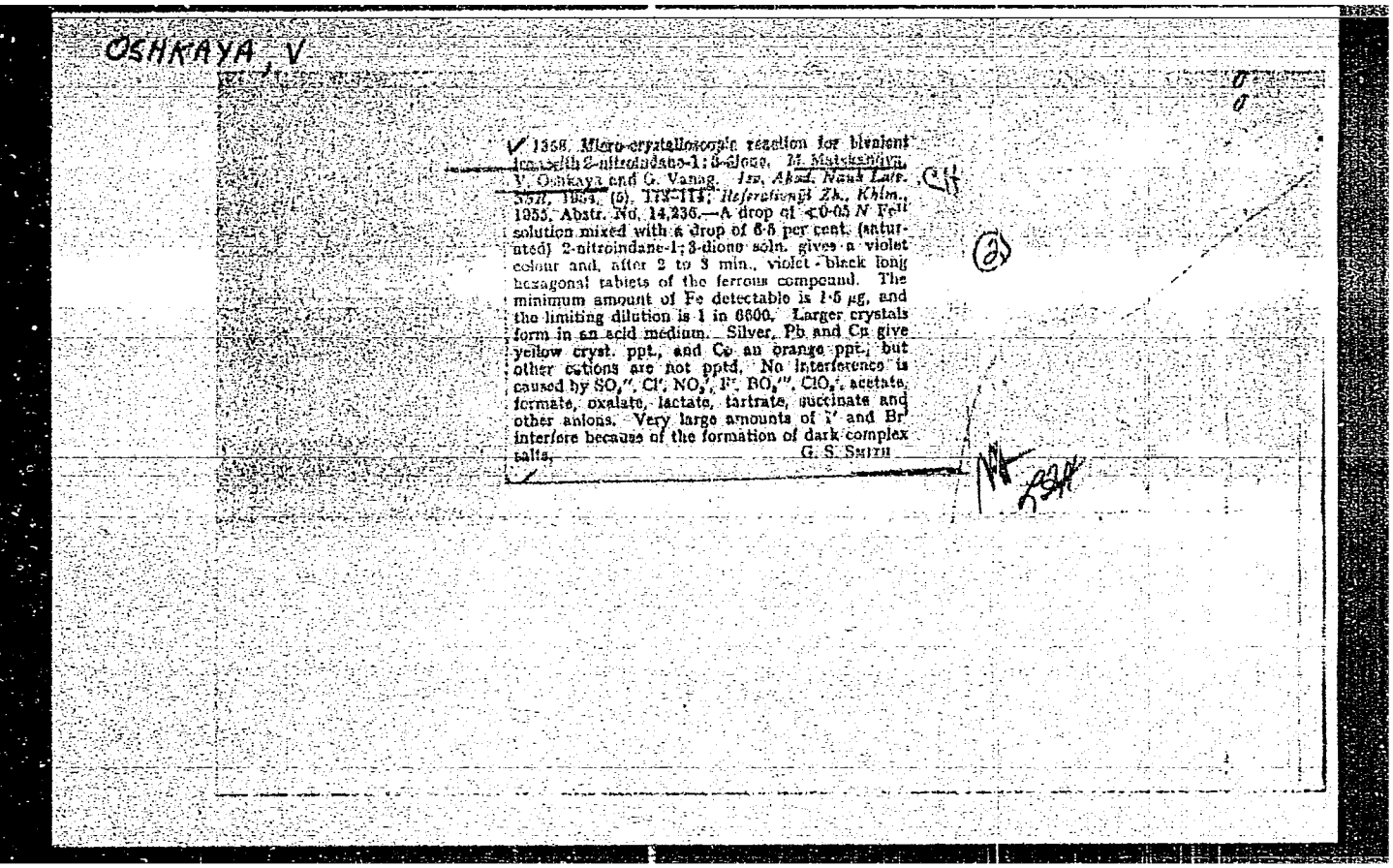
Potentials of mine ventilation. Gor.zhur. no.3:26-30 Mr '65.
(MIRA 18:5)

1. Glavnyy inzh. Upravleniya gornodobyvayushchey promyshlennosti Pridneprovskogo soveta narodnogo khozyaystva (for Shkuta).

OSHURKOVA, M.V.

First f nd of Heraticae in the Carboniferous sediments of the Karaganda Basin. Dokl. AN SSSR 140 no.[illegible] [illegible].

1. Vsesoyuznyy nauchno-issledovatel'skiy [illegible] Institut.
Submitted July 1[illegible], [illegible].



OSHKAYA, V

1358. *Micro-crystalloscopic reaction for bivalent iron with 2-nitroindane-1:3-dione.* M. Matiskiewicz, V. Oshkaya and G. Vanag. *Izv. Akad. Nauk Latv. SSR*, 1954, (5). 111–114; *Referativnyi Zh., Khim.*, 1955, Abstr. No. 14,236.—A drop of <0·05 *N* Fe^{II} solution mixed with a drop of 0·5 per cent. (saturated) 2-nitroindane-1:3-dione soln. gives a violet colour and, after 2 to 3 min., violet-black long hexagonal tablets of the ferrous compound. The minimum amount of Fe detectable is 1·5 μg, and the limiting dilution is 1 in 6600. Larger crystals form in an acid medium. Silver, Pb and Cu give yellow cryst. ppt., and Co an orange ppt., but other cations are not ppted. No interference is caused by SO_4'', Cl', NO_3', F', BO_3''', ClO_4', acetate, formate, oxalate, lactate, tartrate, succinate and other anions. Very large amounts of I' and Br' interfere because of the formation of dark complex salts.

G. S. SMITH

AUTHORS: Vanag, G. Ya., Oshkaya, V. P. SOV/79-28-6-18/63

TITLE: 4-Nitroindandione-1,3(4-Nitroindandion - 1,3)

PERIODICAL: Zhurnal obshchey khimii, 1958, Vol. 28, Nr 6, pp. 1520 - 1524 (USSR)

ABSTRACT: Among the indandione-1,3-derivatives physiologically active compounds as well as reagents of great analytical value were found the latter remark applying especially to the analogs of the indandione. In looking for new analytical reagents in this group the authors believe that, different from those already known, various functional groups can be introduced to the benzene ring of the indandione which fact has hitherto been little used. Based on theoretical considerations the indandione derivatives with a negative (electrophilic)substituent in the benzene nucleus might be of special interest, which also was the reason for preferring the synthesis of 4-nitroindandione-1,3 to the others. Of the two known syntheses of indandione-1,3 only the one could be taken into account which consists of the condensation of the phthalic anhydride with acetic anhydride and the isomerization of the obtained phthalic acetic acid by sodium ethylate to the indandione (see scheme 2). In the present case, however, the acetic anhydride

Card 1/2

had to be replaced by malonic acid and the condensation had to be carried out in pyridine in the presence of some drops of piperidine at low temperature in order to obtain the nitrophthalic acetic acid (see scheme 3). The isomers which had to be theoretically expected did not have to be determined as they had to yield one and the same product in the regrouping, viz. the 4-nitroindandione-1,3 (IV), which was proved by the experiment. This final product does not crystalize well, it dissolves with an orange color in alkali liquor and in acidification is again separated as initial product. On heating with water the 4-nitroindandione-1,3 turns red-violet. The dioxime (V) proves that it contains two carbonyl groups. The product of the conversion with benzaldehyde (VI) points to the presence of an active methylene group. Thus the formula (IV) set up for the 4-nitroindandione-1,3 does not cause any contradictions. There are 6 references, 3 of which are Soviet.

ASSOCIATION: Latviyskiy gosudarstvennyy universitet i Rizhskiy khimikofarmatsevticheskiy zavod (Latvian State University and Riga Chemical-Pharmaceutical Factory)

SUBMITTED: June 10, 1957

OSHKAYA, V.[Oskaja, V.] (Riga); VANAG, G.[Vanags, G.](Riga)

Condensation of phthalic anhydride and phenylacetic acid in triethylamine and acetic anhydride solution. Vestis Latv ak no.3: 67-76 '61. (KRAI 10:9)

1. Akademiya nauk Latviyskoy SSR, Institut organicheskogo sinteza.

(Phthalic anhydride) (Phenylacetic acid)
(Triethylamine) (Acetic anhydride)

OSHKAYA, V.[Oskaja, V.]; VANAG, G.[Vanags, G.]

Condensation of nitrophthalic anhydrides with phenyl acetic acid in triethylamine and acetic anhydride solution. Vestis Latv ak no.6: 57-64 '61.

1. Akademiya nauk Latviyskoy SSR, Institut organicheskogo sinteza.

(Acetic acid) (Phthalic anhydride)

OSHKAYA, V.[Oskaja, V.]; VANAG, G.[Vanags, G.]

Condensation of phthalic anhydride with benzyl cyanide and methyl-bezyl ketone in a solution of triethylamine and acetic anhydride. Vestis Latv ak no.8:45-52 '61.

1. Akademiya nauk Latviyskoy SSR, Institut organicheskogo sinteza.

OSHKIN, N. F., inzh.

Effective technique in repairing the fuel conduit of the 1D12 diesel. Elek. i tepl. tiaga 6 no.9:27 S '62. (MIRA 15:10)

1. Depo Tallin-Vyayke Estonskoy dorogi.

(Diesel locomotives—Repairs)

OSHKIN, N.F.

Repair of a preheater boiler of the TU2 diesel locomotive. Elek. i tepl.tiaga 7 no.2:7 F '63. (MIRA 16:2)

1. Nachal'nik depo Tyuri Pribaltiyskoy dorogi.
(Diesel locomotives)

OSHKIN, P. A.

BLYUMENTAL', R.M.; GIRICH, A.I.; GONCHARIK, A.K.; GUSEVA, T.P.; ZHITKOVA, L.A.; IOFFE, A.M.; KULEMIN, P.D.; LEVINA, L.I.; OSHKIN, P.A.; PAPROTSKIY, T.V.; RYAKHINOV, A.N.; SAMSONOV, N.A.; TULAYKOV, V.N.; USTINOV, I.M.; FAYN, B.P.; SHIFRIN, D.L.; KOLOTILOV, Vasiliy Ivanovich, red.; SVYATITSKAYA, K.P., vedushchiy red.; TROFIMOV, A.V., tekhn.red.

[Equipment for the petroleum industry] Neftianoe oborudovanie. Vol.5 [Petroleum valves and fittings] Neftianaia armatura. Moskva, Gos. nauchno-tekhn.izd-vo neft. i gorno-toplivnoi lit-ry. 1958. 247 p. (MIRA 12:1)

(Petroleum industry--Equipment and supplies)

OSHKIN, P.A., inzh.

Shortcomings of E-505 excavators. Mekh. stroi. 17 no.10:23-24 O '60. (MIRA 13:10)
(Excavating machinery)

STARSHINOV, B.N.; OSTROUKHOV, M.Ya.; KOCHINEV, Ye.V.; Prinimali uchastiye: TARASOV, D.A.; SOROKA, P.F.; KARACHENTSEV, M.D.; OS'KIN, V.T.; KORNEV, V.K.; POPOV, Yu.A.; DOLMATOV, V.A.; AYUKOV, A.S.

Blowing-in of large blast furnaces. Sbor.trud. UNIIM no.11:27-32 '65. (MIRA 18:11)

OSHKINA, N.I.; SKLOVSKIY, I.V., red.; NIKITENKO, A.A., vedushchiy red.; POLOSINA, A.S., tekhn.red.

[Catalog; Spare parts for petroleum equipment] Katalog; Zapasnye chasti k neftianomu oborudovaniiu. Moskva, Gos. nauchno-tekhn. izd-vo neft. i gorno-toplivnoi lit-ry. Pt.2. [Equipment for drilling wells] Oborudovanie dlia bureniia skvazhin. Section 5. [Pulley blocks] Talevye bloki. No.1. [U4-130-3 pulley blocks] Talevyi blok U4-130-3. 1956. 6 p. (MIRA 11:5)

1. Soyuzneftemburmashremont, Gosudarstvennyy, soyuznyy trest. (Pulleys)

OSHKINA, N.I.; KATS, I.N.; PONOMAREVA, Ye.V.; SKLOVSKIY, I.V., red.; PETROVA, Ye.A., red.; KHLEBNIKOVA, L.A., tekhn.red.

[Catalog of spare parts for petroleum equipment] Katalog: Zapasnye chasti k neftianomy oborudoveniiu. Moskva, Gos. nauchno-tekhn.izd-vo neft.i gorno-toplivnoi lit-ry. Pt.2. [Equipment for drilling wells] Oborudovanie dlia bureniia skvazhin. Section 17. [Stationary drilling installations] Ustanovki burovye statsionarnye. No.1. [Uralmash 5D drilling rig with five diesel drive] Burovaia ustanovka Uralmash 5D piatidizel'nyi privod. 1957. 71 p. (MIRA 11:1)

1. Soyuzneftburmashremont, Gosudarstvennyy soyuznyy trest. (Oil well drilling--Equipment and supplies)

OSHKINIS, B

AID P - 731

Subject : USSR/Aeronautics

Card 1/1 Pub. 58 - 12/16

Author : Oshkinis, B., Designer

Title : Glider BRO-9

Periodical : Kryl. Rod., 10, 18-19, O 1954

Abstract : This is a description and diagrams of the glider BRO-9. designed by the author in 1952. Diagrams.

Institution : Written in the city of Kannas

Submitted : No date

OSHKINIS, B.

AID P - 1650

Subject : USSR/Aeronautics

Card 1/1 Pub. 58 - 9/19

Author : Oshkinis, B., Designer

Title : Glider "Pioner" (Article in the series "What mass-produced glider do we need?")

Periodical : Kryl. rod., 3, 9-10, Mr 1955

Abstract : The author suggests the mass production of a simple glider for primary training. He gives its specifications, performances, and building cost. Photos, diagram.

Institution: DOSAAF (All-Union Voluntary Society for the Promotion of the Army, Aviation and the Navy)

Submitted : No date

OSHKINIS, B

AID P - 5543

Subject : USSR/Aeronautics - Training of Glider Pilots

Card 1/1 Pub. 58 - 2/20

Author : Oshkinis, B., Designer, Head of the Kaunasskaya (Lithuanian SSR) Glider Station, DOSAAF.

Title : A new contrivance for ground training of glider pilots

Periodical : Kryl. rod., 1, 2-5, Ja 1957

Abstract : Detailed description of a contrivance permitting a thorough ground training of the future glider pilots. The aircraft serving as a trainer is placed on a long counter-balanced beam, which is turned in the direction of the wind. 3 drawings, 2 photos.

Institution : None

Submitted : No date

OSHKINIS, Bronyus [Oškinis, B.]

Perfect organization brings a good flying time. Kryl.rod. 1[illegible] no.9:14 S '61. (MIRA 14:9)

1. Nachal'nik Kaunasskoy planernoy stantsii.
(Kaunas---Gliding and soaring)

KHAMZIN, R.G.; VASIL'YEV, I.P.; OSHITKO, V.M.

Exploitation of nonuniform producing layers of horizon D_1 in the Zay-Karatay area of the Romashkino oil field. Geol. nefti i gaza 9 no.4:10-13 Ap '65. (MIRA 18:8)

1. Leninogorskneft'.

POLAND/Cultivated Plants - Ornamental. M.

Abs Jour : Ref Zhur - Biol., No 4, 1958, 15906

Author : W. Oshkinis
Inst : -
Title : Comparative Research on the Effect of Several Growth Substances on the Graft Rooting of the William Sims Variety of Hothouse Carnation, Dianthus caryophyllus. (Sravnitel'nyye issledovaniya vliyaniya nekotorykh rostovykh veshchestv na ukoreneniye cherenkov teplichnoy gvozdiki Dianthus caryophyllus sorta Vil'yam Sim).

Orig Pub : Acta agrobot., 1956, 4, 191-202.

Abstract : The best results were obtained from naphthyl acetic acid in 0.001% solution, this process not only yielding a higher percentage of root taking but also speeding it up.

Card 1/1

L 42005-65 EWT(m)/EPF(c)/EWA(d)/EPR/T/EWP(t)/EWP(z)/EWP(b)/EWA(c) Pad/
Ps-4 IJP(c) JD/HW/WB/GS

ACCESSION NR: AT5009565 UR/0000/64/000/000/0147/0153 37 24+1

AUTHOR: Arkharov, V.I. (Professor); Gerasimov, A.F.; Oshkova, T.V.; Kudinova, V.A.; Puzanova, L.V.

TITLE: A study of the high-temperature oxidation of copper-nickel alloys

SOURCE: AN SSSR. Institut fizicheskoy khimii. Mekhanizm vzaimodeystviya metallov s gazami (Mechanism of interaction of metals and gases). Moscow, Izd-vo Nauka, 1964, 147-153

TOPIC TAGS: copper alloy, nickel alloy, copper oxidation, nickel oxidation, high temperature oxidation, xray analysis, oxide scale, oxidation kinetics

ABSTRACT: This work is a continuation of studies on the physical mechanism and kinetics of reaction diffusion in metal-gas systems. It deals with the oxidation of Cu-Al alloys [illegible] temperatures between 700 and 1200C. The following

parameter of ... and metallographic ...
weight of samples during oxidation), and metallographic ...

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L 42005-65

ACCESSION NR: AT5009565

nickel were also oxidized for comparison. On the basis of the experimental data obtained, an interpretation of the oxidation mechanism is given in terms of three cases: (1) alloys with minute amounts of copper, (2) alloys with appreciable amounts of copper (about 25%), and (3) alloys with over 60% copper. Orig. art. has: 4 tables and 4 formulas.

ASSOCIATION: None

SUBMITTED: 26Oct64 ENCL: 00 SUB CODE: MM

NO REF SOV: 005 OTHER: 006

Card 2/2

RUMANIA / Cultivated Plants. Fruits, Berries. M-7

Abs Jour : Ref Zhur - Biologiya, No 13, 1958, No. 58764

Author : Konstantinescu, G.; Oshlobyanu, M.
Inst : Rumanian Agronomical Institute
Title : Theories Concerning Green Pruning-Pinching and Nipping of the Vine

Orig Pub : Biol. Zh. Akad. RNR, 1956, 1, No 2, 187-194

Abstract : Observations conducted by the Scientific Research Agronomical Institute in various regions of RPR in 1949-1953 showed that the appearance of embryos of racemes on the sprout does not cause an intensive consumption of assimilation products before blooming. The growth slows up with the beginning of blooming. The greatest amount of withering away of buds takes place during the budding phase or at the beginning of the phase when the flowers open. Nipping, carried out

Card 1/3

Abs Jour : Ref Zhur - Biologiya, No 13, 1958, No. 58764

before blooming, interrupts the growth for 7-10 days and intensifies the flow of nutritious substances to the racemes, thus diminishing the number of unsuccessful buds. The yield of the Galbene and of Furmint varieties increased by 21-32%. Nipping also increases the yield during the following year. The berries consume only 1/3-1/4 of nutritious substances during the ripening period. After ripening of the berries, the growth of the sprouts continues during 10-15 days. Pinching carried out on 6 leaves from the top of the sprout during the ripening period, or 2-3 days before this period, removes the "consuming" leaves and arrests the growth which would cause the appearance of such leaves. The racemes, in this case, receive better nutrition during 1-2 months. For varieties growing with minimum daylight (Proles orientalis, Subproles

Abs Jour : Ref Zhur - Biologiya, No 13, 1958, No. 58764

antasiatica Negrul), pinching is an absolute necessity. The yield increment in this case is 9.1-9.4%. --
I. K. Fortunatov

RUMANIA/Cultivated Plants - Fruits. Berries. M.

Abs Jour : Ref Zhur - Biol., No 4, 1958, 15783

Author : G. Konstantinesku, M. Oshlobyanu
Inst : -
Title : The Grafting and Pinching Out of Grape Vines. (Prishchipyvaniye i chekanka vinogradnoy lozy).

Orig Pub : Gradina, via si livada, 1956, 5, No 5, 35-38.

Abstract : Tests were conducted in 1949-1953 at 6 viticultural centers. The grafting was done before florescence, during the flowering time, on 2-3 days before the berries ripening began and in the period of termination of shoot growth with 2, 3, 6 and 8 leaves. The pinching out was performed in the middle of the flowering period on the second-third day after the flowers dropped off, up to the period of berry ripening and in the period of shoot growth termination, with 5, 10 and 15 leaves. It is recommended that shoot grafting be done

Card 1/2

M

RUMANIA/Cultivated Plants. Fruits. Berries.

Abs Jour: Ref Zhur-Biol., No 5, 1958, 20520.

Author : M. Oshlobyanu
Inst : Not given.
Title : The State of Viticulture in the People's Republic of Bulgaria. (Sostoyaniye vinogradarstva v Narodnoy Respublike Bolgarii).

Orig Pub: An. Rom.-Sov. Ser. agric., 1956, 10, No 6, 71-82.

Abstract: Resulting from the author's visit to viticultural and wine-making institutions, organizations and farms in the rayons of the cities of Sofia, Plovdiv, Plevna, Trnovo, Ruse and Stalin, information is given on the state of wine making and viticulture, the area taken up by vineyards, the correspondences between table and wine varieties, a list of the most prevalent varieties of grapes and species

Card : 1/2

SANDLER, A.B.; OSHLYANSKIY, A.I., inzh.

Manufacture of women's footwear with uppers made from artificial leather. Kozh.-obuv. prom. 7 no. 11:33-38 N '65 (MIRA 19:1)

1. Glavnyy inzhener Belotserkovskoy obuvnoy fabriki No. 14 (for Sandler).

137-58-6-11954

Translation from Referativnyy zhurnal, Metallurgiya, 1958, Nr 6, p 110 (USSR)

AUTHORS Ryabov, Yu.F., Okunev, A.I., Kirr, L.D., Oshman, V.A.

TITLE Distribution of Certain Rare and Disseminated Elements in the Treatment Products of Copper Ores and Concentrates (Raspredeleniye nekotorykh redkikh i rasseyannykh elementov v produktakh pererabotki mednykh rud i kontsentratov)

PERIODICAL Byul. tsvetn. metallurgii, 1957, Nr 22, pp 24-27

ABSTRACT Tables of the distribution of rare and disseminated elements at various stages of conversion at the Karabash and Krasnoural'sk copper smelter are presented. Under conditions of pyrometallurgical treatment, In, Ge, and Tl chiefly go into the slags, Se and Te go into the blister Cu and the dust, and Cd into the dust.

G.S.

1. Copper ores--Processing 2. Rare earth elements--Determination

Card 1/1

137-58-1-1532

Translation from: Referativnyy zhurnal, Metallurgiya, 1958, Nr 1, p 10 (USSR)

AUTHORS: Okunev, A.I., Kirr, L.D., Oshman, V.A., Ryabov, Yu.F.

TITLE: The Distribution of Rare and Disseminated Elements in the Milling of Ural Copper-and-zinc Ores by Separation of Independent Concentrates (Raspredeleniye redkikh i rasseyannykh elementov pri obogashchenii ural'skikh medno-tsinkovykh rud s vydeleniyem samostoyatel'nykh kontsentratov)

PERIODICAL: Byul. tsvetn. metallurgii, 1957, Nr 23, pp 12-13

ABSTRACT: The Unipromed' Institute has made a study of the distribution of rare and disseminated elements among the various milling products at sections of the Krasnoural'sk and Karabash Ore Milling plants. The results of the analyses show that the Cd, In, and Ge contents of the Zn concentrate exceed many times over the contents thereof in the starting specimen. However, the total extraction thereof in the Zn concentrate is comparatively low, and it is 67-85% in the pyrite concentrate and tailings. The concentration of rare elements in the Cu concentrate is less, owing to the separation of Zn concentrate, than is the case in flotation involving a combined Cu-Zn concentrate.

A.Sh.

Card 1/1

1. Copper ores--Processing 2. Zinc ores--Processing
3. Rare earth elements--Availability

5(2)

AUTHORS: Oshmyan, V. A., Candidate of Chemical Sciences, Ogorodnikov, K. V. SOV/32-24-11-34/37

TITLE: Seminar on the Analytical Chemistry of Scattered Elements (Seminar po analiticheskoy khimii rasseyannykh elementov)

PERIODICAL: Zavodskaya Laboratoriya, 1958, Vol 24, Nr 11, pp 1423-1423 (USSR)

ABSTRACT: The seminar mentioned in the title was held by the Sverdlovskoye otdeleniye Vsesoyuznogo khimicheskogo obshchestva im. D.I. Mendeleyeva (Sverdlovsk Branch of the All-Union Chemical Society im. D.I. Mendeleyev), the Ural'skiy Dom tekhniki (Ural House of Technology), and other organizations at the city of Sverdlovsk from June 23 to 28, 1958. The seminar was attended by 50 participants from 27 plants, mines, and research institutes of various fields. In the plenary sessions reports were given on the use and modern methods of determination of rare and scattered elements. The participants in the seminar dealt with practical new analytical methods in this field. The participants were informed about rapid methods. In accordance with the desire expressed by the participants, the Ural House of Technology

Card 1/2

Seminar on the Analytical Chemistry of Scattered Elements SOV/32-24-11-34/37

will hold a similar seminar on rare alkaline earth elements in 1958.

Card 2/2

S/137/62.0[illegible]
A160/A101

AUTHOR: Osmar., V. A.; Chistyakova, A. P.

TITLE: Use of titanic acid as the collector, when preci[illegible] amounts of indium

PERIODICAL: Referativnyy zhurnal, metallurgiya, no. 3, 1962, 2, [illegible] ("Khim., fiz.-khim. i spektr. metody issled. rud rock[illegible] elementov", Moscow., Gosgeoltekhizdat, 1961, 95 - 10[illegible]

TEXT: A polarographic method of determining In after its prel[illegible] ration, in the form of a basic sulfite, by a combined hydrolysis with [illegible] is described. An amount of 1 - 10 g ore (a concentrate, matte, dust, [illegible] tion, etc.), containing 0.01 - 0.0002 % of In, is heated together with HCl, evaporated down to a syrup-state, supplemented with 10 - 50 ml of aqua regia (HCl is added too, if necessary), and boiled down to dryness. Slags are decomposed in the presence of 1 - 5 ml HF. If Sn is present, it is removed by a twice-repeated treatment with HBr. The As is removed by a single evaporation with HCl in the presence of hydrazine hydrochloride and KBr. The residue is dissolved in

Card 1/3

S/137/62/[illegible]
A160/A[illegible]

Use of titanic acid as ....

100 - 250 ml hot water, after which one adds 1 - 5 g hydrazine [illegible]
5 ml of 3.2 % solution of $(NH_4)_2MoO_4$, and neutralizes the corpos[illegible]
until the precipitation of $Fe(OH)_3$, which is followed by a [illegible]
and filtering-out the insoluble radical, containing $PbSO_4$, [illegible]
sulfate, and Cu. The filtrate is neutralized with NH_4OH until the [illegible]
greenish hydrates become only slowly soluble (pH 3). Then the filtrate is [illegible]
plemented with 3 g NH_4Cl, 5 g $Na_2S_2O_5$ and boiled, as it amounts [illegible]
for 15 - 20 minutes until the SO2 odour almost completely vanish[illegible]
with bromophenol blue; blue coloring at pH = 4 - 4.5). At this [illegible]
duces 10 ml of 0.1 % solution of TiO_2, diluted in advance to [illegible]
in the presence of 1 drop of H_2O_2 and exactly neutralized by [illegible]
the disappearance of orange color. Immediately afterwards the [illegible]
c. . % Na_2HPO_4 solution, and boils the composition for 15 more [illegible]
deposit coagulates. The next day the deposit is filtered, washed [illegible]
solution of NH_4Cl, and treated together with the filter, while [illegible]
heated, with 30 - 50 ml HCl (1 : 1) and 5 ml H_2O_2. The paper is [illegible] a [illegible]
washed with HCl. The filtrate is precipitated with the aid of NH_4OH and filter[illegible]

Card 2/3

S/137/62/000/003/175/[illegible]
A160/A101

Use of titanic acid as ....

The precipitate is dissolved in HCl (1 : 1), the solution is boiled down to an amount of 5 - 10 ml and then transferred into a 25 ml flask, filling the latter up to mark with HCl (1 : 1), treated with reduced Fe powder (Sb, Bi and Hg separate themselves on Fe powder and do not hinder the determination of In). The solution is filtered and In is polarographically analyzed, the rheochord voltage being 0.9 v. This method should not be used for analyzing rich Pb, Mo and W concentrates.

N. Gertseva.

[Abstracter's note: Complete translation]

Card 3/3

OSHMAN, V.A.; CHISTYAKOVA, A.P.

Polarographic determination of cadmium on an acid sulfate-iodide background. Zav.lab. 27 no.5:532-536 '61. (MIRA 14:5)
(Cadmium—Analysis) (Cadmium iodide)
(Sulfuric acid)

OSHMAN, V.A.; VOLKOV, V.M.

Simplified method of determining germanium with phenylfluorone. Zav.lab. 27 no.11:1341-1343 '61. (MIRA 14:10)

1. Ural'skiy nauchno-issledovatel'skiy i proyektnyy institut mednoy promyshlennosti.
(Germanium—Analysis)

ACCESSION NR: AR4015685 S/0081/63/000/023/0129/0130

SOURCE: RZh. Khimiya, Abs. 23G72

AUTHOR: Oshman, V. A.; Chistyakova, A. P.

TITLE: Polarographic determination of indium after its isolation as a basic sulfite

CITED SOURCE: Tr. Ural'skogo n.-i. i proyektn. in-ta medn. prom-sti, vy*p. 6, 1962, 245-250

TOPIC TAGS: polarography, indium, indium sulfite, indium analysis

TRANSLATION: A polarographic method was developed for the determination of In in concentrations of 0.1-0.0002% in ores, concentrates, matte, powders, sublimates, etc. It has been found that during the combined hydrolysis of In with 10 mg of TiO_2 in 200 ml solution containing 5g $Na_2S_2O_5$ and 100 mg Na_2HPO_4, complete precipitation of In as the basic sulfite is achieved at pH_3. The hydrolysis of In and Ti must be synchronized; that is, the Ti must go into solution after removal of the major portion of the SO_2 by boiling. The introduction of Ti to $Na_2S_2O_5$ leads to its premature hydrolysis and the

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ACCESSION NR: AR4015685

coprecipitation of In is not quantitative. For the determination of In, 1-10 g of sample is decomposed by any acidic method and the solution evaporated to dryness. Sn, if present, is removed by treating the dried residue with $HBr-Br_2$ mixture. After evaporation, the residue is dissolved in hot water (100-250 ml), 1-5 g $N_2H_4 \cdot H_2SO_4$ and 5 ml 0.2% ammonium molybdate solution are added, the mixture is neutralized with NH_4OH until $Fe(OH)_3$ begins to separate, boiled 10-15 minutes and filtered. The insoluble precipitate contains $PbSO_4$, Te, the double sulfate of hydrazine and Cu. The filtrate is neutralized with NH_4OH to pH 3, 5g NH_4Cl and 5g $Na_2S_2O_5$ are added, followed by dilution to 250-400 ml. The solution is boiled to disappearance of the SO_2 odor and the formation of a Feoxide film on the surface of the solution. Meanwhile, a 0.1% TiO_2 solution was diluted to to 100 ml after addition of 1 drop H_2O_2, cooled and carefully neutralized with NH_4OH to the disappearance of the color of pertitanic acid. 10 ml of this solution is added to the boiled solution immediately after the disappearance of the SO_2 odor. Then 10 ml of 1% Na_2HPO4 is added and the solution is boiled until the precipitate coagulates. On the following day the precipitate is filtered out and washed with a 2% solution of NH_4Cl. The filter with the precipitate is treated with 30-50 ml HCl (1:1) and 5 ml H_2O_2, the paper is filtered out and washed with the HCl solution. Ammonia is then introduced into the filtrate, and 1-2 ml of 10% $FeCl_3$ solution may

ACCESSION Nr: AR4015685

be added if necessary. The precipitate is filtered out and dissolved on the filter in HCl (1:1) The solution is concentrated to 5-10 ml, diluted to 25 ml with HCl (1:1) and reduced with powdered metallic Fe which was prepared in a stream of H_2. The solution is filtered and the filtrate is polarographed. The time required for the determination of In by this method is only 1/2 - 2/3 as long as by the Gintsvetmet method (cementation of In on Zn dust).
N. Gertseva

SUB CODE: IC DATE ACQ: 09Jan64 ENCL: 00

Card 3/3

VOL'KENSHTEYN, A.A., kand.tekhn.nauk; OSHMARIN, A.G., inzh.

VFM-57 visual laboratory photometer (for low brightness). Svetotekhnika 4 no.12:19-21 D' 58. (MIRA 11:12)
(Photometers)

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Transition to parasitism by ancestors of helminths caused reduction or elimination of some organs and development of others. Some unusual filariae, parasitic in jays in the Far East, have lost all but a vestige of the intestine attached to the esophagus. Intestinal functions are performed by the esophagus and, perhaps, the whole surface of the nematode's body. Describes species in detail. Submitted by Acad K. I. Skryabin 9 Apr 49.

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9. Monthly List of Russian Accessions, Library of Congress, ~~September~~ 1952, Uncl.

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(Maritime Territory--Worms, Intestinal and parasitic)

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USSR/Zooparasitology - Parasitic Worms.

Abs Jour : Ref Zhur - Biol., No 5, 1958, 19999

Author : Oshmarin, P.G.
Inst : -
Title : Tetrameridae (Spirurata, Tetrameridae) of Domestic and Wild Fowl in the Maritime Region.

Orig Pub : Tr. Dalnevost. fil. AN SSSR, ser. zool., 1956 3 (6), 281-314

Abstract : This is a list of tetrameridae species found in fowl of the maritime region, with an indication of their hosts, distribution within the region, the extensiveness and intensity of infection, as well as the periods of helminth detection. A description and illustration of new species: Tetrameres (Tetramares) para-araliensis from lapwings, T. (Petrowimere) striatus from ducks, T. (Gynaecophila) schigini from auburn herons, Microtetrameres asymmetricus from speckled magpies, a sub-species of M. helix asiaticus from

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(MIRA 11:9)

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(Parasites--Cattle)

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[Helminth fauna of Far Eastern salmonids in connection with the problem of local stocks and migration routes of these fishes] Gel'mintofauna dal'nevostochnykh lososevykh v sviazi s voprosom o lokal'nykh stadakh i putiakh migratsii etikh ryb. Vladivostok, Primorskoe knizhnoe izd-vo, 1959. 72 p. (MIRA 13:10)
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(Pastures and meadows)

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Morphology and features of the concentration of Filaroides martis in the kolinsky of the Maritime Territory. Soob.DVFAN SSSR no.11: 152-155 '59. (MIRA 13:11)

1. Dal'nevostochnyy filial imeni V.L.Komarova Sibirskogo otdeleniya AN SSSR.

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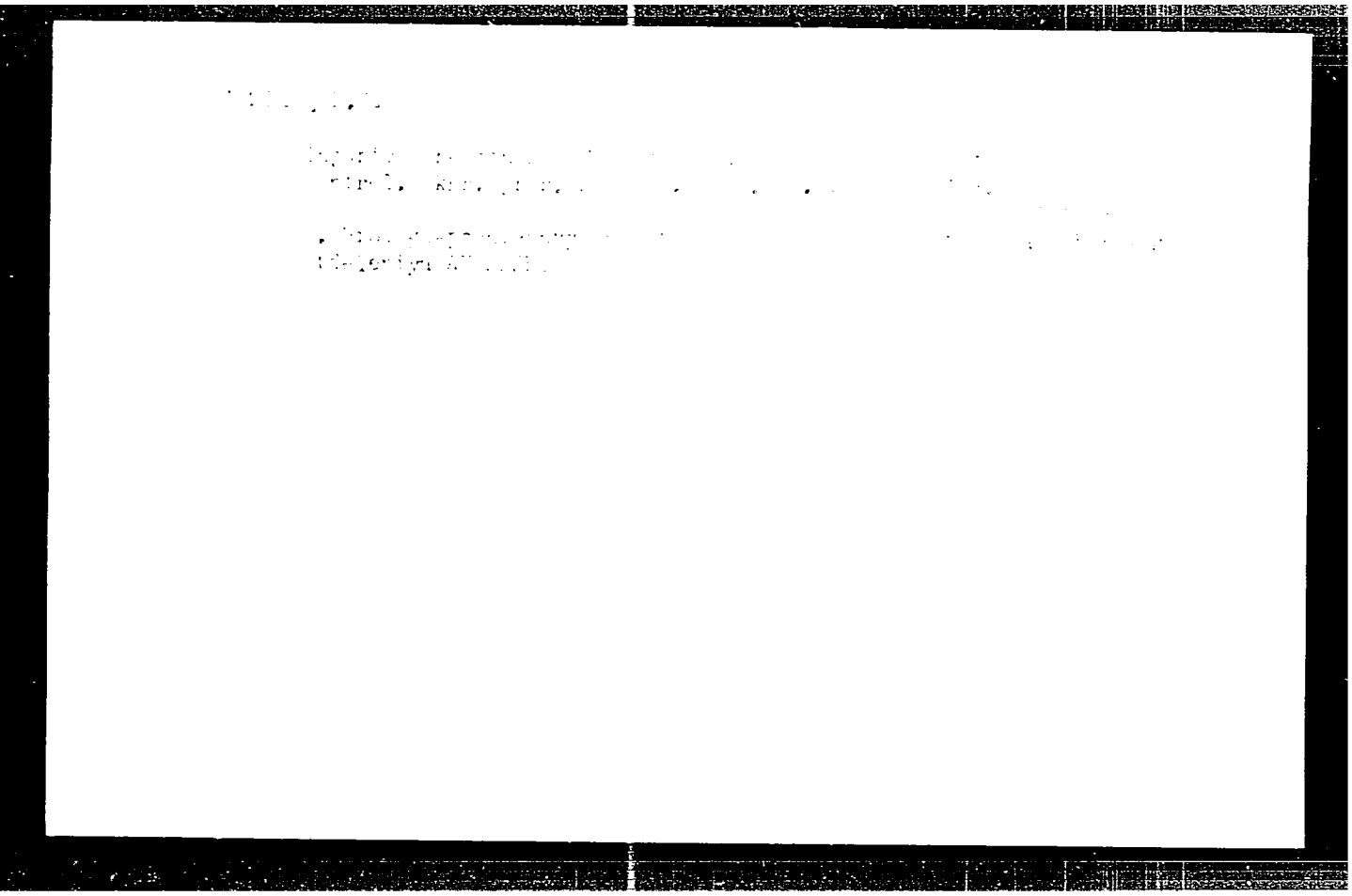

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