

OGLOBLINA, Z.B.

[Troianov, Aleksei Alekseevich] Aleksei Alekseevich Troianov. Vest.  
khir. 70 no.1:14-22 '50. (CML 19:1)

1. Of the Hospital Surgical Clinic of the First Leningrad Institute  
Imeni Academician I.P.Pavlov (Director -- Yu.Yu. Dzhanalidze, deceased).

OGLOBLINA, Z. V.

New method of closing biliary fistulas. *Khirurgia*, Moskva.  
no. 10:70-72 Oct. 1950. (CML 20:1)

1. Of the Hospital Surgical Clinic (Director -- Yu. Yu. Dzhanelidze), First Leningrad Medical Institute imeni Academician I. P. Pavlov. 2. Z. V. Ogloblina has the title of Professor.

OGLOBLINA, Z. V.

Ogloblina, Z. V. "Acute arterial impassibility," Trudy XXV Vsesoyuz. s'yezda khirurgov. Moscow, 1948, p. 536-42

SO: U-3244, 10 April 1953, (Letopis 'nykh Statel', No. 3, 1949

143905-66 EWT(m)/EWP(t)/ETT IJP(c) JD  
ACC NR: AP6015623 (A) SOURCE CODE: UR/0413/66/000/009/0017/0017

INVENTOR: Smorodinnikov, A. V. ; Khokhlov, D. G. ; Shamarin, V. A. ;  
Ogloblina, Z. A.

30  
B

ORG: none

TITLE: Method of obtaining phosphorus. Class 12, No. 181062 [announced by  
the Ural Scientific Research and Planing Institute for Concentration and Mechanical  
Processing of Minerals (Ural' skiy nauchno-issledovatel' skiy proyektnyy institut  
po obogashcheniyu i mekhanicheskoy obrabotke poleznykh iskopayemykh)]

SOURCE: Izobreteniya, promyshlennyye obraztsy, tovarnyye znaki, no. 9, 1966, 17

TOPIC TAGS: phosphorus, phosphate mineral, sintering, fluxing additive

ABSTRACT: An Author Certificate has been issued for a method of obtaining a  
phosphorus from raw phosphate in the presence of fluxing additives by sintering  
phosphorite with subsequent electrothermal reduction of the agglomerate obtained.  
To simplify the process, the fluxing additives are introduced before sintering the  
phosphorite. [Translation] [INT]

SUB CODE: 11/ SUBM DATE: 03Dec64/

Card 1/1 *gjm*

UDC: 661.631.3

OGLOBLINA, Yelena Fedorovna; KUCHINSKIY, V., red.; POLTORAK, I., tekhn.red.

[Weed control in cotton and alfalfa fields] Bor'ba s sorniakami  
na khlopkovykh i liutsernovykh poliakh. Stalinabad, Tadzhikskoe  
gos. izd-vo, 1958. 9 p. (MIRA 12:1)  
(Cotton growing) (Alfalfa) (Weed control)

OGIOBLINA, Ye.F.

Periods and methods of disbudding in the Gissar Valley and in districts of Leninabad Province. Izv.Otd.est.nauk AN Tadzh. SSR no.14:77-84 '56. (MLRA 9:10)

1. Otdel khlopkovodstva AN Tadzhikskoy SSR.  
(Gissar Valley--Cotton growing)  
(Leninabad Province--Cotton growing)

URYVAYEVA, G.D.; PENDYURINA, T.Ye.; OGLOBLINA, T.Ye.

Effect of lime and gypsum on the process of ~~hardening~~ of tetra-  
calcium aluminoferrite. Izv. SO AN SSSR no.11 Ser.khim.nauk no.  
3:21-25 '63. (MIRA 17:3)

1. Khimiko-metallurgicheskiy institut Sibirskogo otdeleniya AN SSSR,  
Novosibirsk.

STRUVE, M.Ye. (Moskva, V-415, Leninskiy prospekt, d.104, kv.58); ROSKIN, G.I.;  
OGLOBLINA, T.A.

Cytochemically distinct types of human tumor cells. Vop onk. 10  
no.8:41-46 '64. (MIRA 18:3)

1. Iz laboratorii eksperimental'noy tsitologii i ts'tokhimii  
rakovoy kletki, kafedry tsitologii i gistologii (zav. - prof. G.I.  
Roskin) Moskovskogo gosudarstvennogo universiteta imeni Lomonosova.



VOL' POVA, Ye.G.; OGLOBLINA, L.I.

Studying the conditions governing the regeneration of a silicotungstic catalyst for the polymerization of propylene and the economics of the process. Trudy GrozNII no. 15:265-270 '63. (MIRA 17:5)

BRESHCHENKO, Ye.M.; OGLOBLINA, L.I.; EMMANUILOVA, Ye.M.

Stabilization of a solution of sodium aluminate by condensed  
sulfite-distiller's grains. Khim. prom. no.5:392-393 My '63.  
(MIRA 16:8)

Polymerization of propylene ...

S/065/63/000/001/002/005  
E075/E436

1000 hours. After 1100 hours the yield decreased to 60% and then gradually to 24%. After activation at 420°C the active life of the catalyst was extended to 2520 hours, the yield of polymers for this time amounting to 440 kg/kg of the catalyst. The production costs of detergent alkylate using the silico-tungstic acid catalyst are about 25% of those produced with a phosphoric acid catalyst. There are 2 figures. ✓

ASSOCIATION: Groz NII

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S/065/63/000/001/002/005  
E075/E436

AUTHORS: Vol'pova, Ye.G., Ogloblina, L.I.  
TITLE: Polymerization of propylene on silico-tungstate  
catalysts  
PERIODICAL: Khimiya i tekhnologiya topliv i masel, no.1, 1963,  
19-24

TEXT: The work was carried out to discover active supports for silico-tungstic acid from minerals available in the Soviet Union. The materials investigated were: Kieselguhr, Askan and Troshkov clays (catalysts), activated silica-alumina and alumina. The polymerization of propylene at 170 to 180°C, 60 atm and 0.3 h<sup>-1</sup> space velocity of the feed proceeded at the most rapid rate when silico-tungstic acid was supported on silica-alumina pellets and Troshkov clay, the yield of polymers being 75 and 65% respectively. The yield was below 25% for all the other supports examined. For Troshkov clay, the polymers contained 14.3% trimers and 34% tetramers, the yield of all the polymers in the first 24 hours being up to 67.7%, but decreasing steadily with time. This did not take place on alumina, the catalyst conserving its activity for

Card 1/2

BRESCHENKO, Ye.N.; OGLOBLINA, L.I.

Rapid method of determining the concentration of silicic acid in the medium of liquid phase. *Trav. Z. O. S. no. 10: 6-10.* (1957)

1. Gromozhkiy neftyanoy nauchno-issledovatel'skiy institut.  
(Gloss) (Silicic acid)

KAMAKIN, N.M.; ROMANKOVA, I.K.; OGLOBLINA, L.I.; NESMEYANOVA, T.S.

Causes of the aging of aluminum silicate catalyst. Trudy GrozNII  
no.4:90-101 '59. (MIRA 12:9)  
(Catalysis) (Aluminum silicates)

contains numerous tables with the characteristics of different petroleum products obtained from refinery processing units. Each article is accompanied by references.

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contains numerous tables with the characteristics of different petroleum products obtained from refinery processing units, plant and petrochemical refinery sections. Each article is accompanied by references.

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OGLOBLNA, L. I.

11 (2, 4) PHASE I BOOK EXPLOITATION 30V/2213

Groznyy. Neftyanoy nauchno-issledovatel'skiy institut

Khimiya i tekhnologiya pererabotki nef'ti i gaza (Chemistry and Technology of Petroleum and Gas Refining Processes) Moscow, Gosstokhkhimizdat, 1959. 278 p. (Series: Izs. Trudy. vyp. 4) 2,500 copies printed.

Executive Ed.: T. D. Yefremov; Tech. Ed.: A. S. Polozina; Editorial Board: A. Z. Dorogochinakiy (Chairman), B. K. Amerkhanov, O. I. Kaz'ain, M. M. Kasakin, V. I. Lavrent'yev, Ye. S. Lvchenko, and M. G. Mitrofanov (Deputy Chairman).

PURPOSE: This book is intended for petroleum engineers and technicians in scientific research institutes, planning organizations, and refineries.

COVERAGE: This collection of technical papers on oil and gas refining were originally discussed at the petroleum refining section of the Third Grozny Scientific-Technical Congress in 1957. The articles have been published to help further the development of the petroleum refining industry and petrochemical industry in the Chechen-Ingush ASSR. The history and significance of the petroleum refining industry in the Grozny region is outlined by A. B. Dorogochinakiy with emphasis on the interdependence of refineries and the aircraft, automobile and rocket manufacturing industries. Change in modern engines decreased use of jet aircraft makes the production of high octane aviation gasoline less important than the production of low octane aviation gasoline. The yield of aviation gasoline requires a quite different refinery run. The yield of aviation gasoline is determined by the presence of a valuable component at the Karabulak-Achaluki fields. The yield of aviation gasoline for manufacturing lubricating oils and paraffin wax, low octane aviation gasoline, and other products is investigated and results of analyses reviewed. The use of the equipment of the fuel producing line of refineries and Grozny has been carried out on the basis of flight tests and pilot plant operations, and a number of reforming and platforming units have been built to upgrade the low octane gasoline produced at Grozny. The authors also conducted to ascertain the advisability of developing the destructive distillation of residues, melt as feed stock. Catalytic beds needed for catalytic cracking units were first put on stream in the Grozny refineries in 1956, and since that time continuous efforts have been made to increase the processing capacity, and improve the regeneration of catalysts. The authors make a number of suggestions to show the throughput of the above units might be increased. The production of different types of pellets and bead catalysts, the contamination of catalysts and their reaction are discussed. The operation of a contact sintering catalyst design, and products yielded by contact sintering catalysts, paraffin and kerosene are discussed. The authors also deal with the design of a way of improving their products. The authors also discuss the dehydration and desulfating of crude oil and petroleum products, and particularly of gases. An extensive number of gas fractionators and compressors were built and installed to produce phenol and acetone from petroleum hydrocarbons. An article on the synthesis of ethyl alcohol from petroleum hydrocarbons and articles on the control and automation of various processes and systems are discussed. The authors also discuss the characteristics of different petroleum products obtained from refinery processing units, pilot plants and petrochemical refinery sections. Each article is accompanied by references.

TABLE OF CONTENTS:

OGLOBLINA, L.I.; YEL'NIKOVSKAYA, N.V.; KAMAKIN, N.M.

Determination of minute quantities of tetraethyllead in benzene solvents. Khim. i tekhn. topl. i masel no.11:72-3 of cover N '57.  
(MIRA 11:1)

1. Groznenskiy nauchno-issledovatel'skiy neftyanoy institut.  
(Gasoline--Analysis) (Lead--Analysis)

A Simple Method of Determining the Activity of an Aluminium Bead  
Catalyst

SOV/65-58-11-14/15

perature (750°C) i.e. a stabilised catalyst). At the same time the activity index was determined according to VTU. Results of experiments on more than 50 samples are given in the form of a graph. Variations in the results obtained by the two methods did not exceed 1.0 - 1.5% and very seldom 2%. The described method can be used in factories producing catalysts for quality control, and also in catalytic cracking plants. The process requires only three hours and one person can carry out two simultaneous experiments; the VTU method is much more time-consuming and requires several operators. There are 1 Table, 1 Figure and 13 References: 8 English and 5 Soviet.

ASSOCIATION: GvozNII

Card 2/2

SOV/65-58-11-14/15

**AUTHORS:** Romankova, I. K; Kamakin, N. M. and Ogloblina, L. I.

**TITLE:** A Simple Method of Determining the Activity of an Aluminium Bead Catalyst (Prostoy metod opredeleniya aktivnosti sharikovo alyumosilikatnogo katalizatora)

**PERIODICAL:** Khimiya i Tekhnologiya Topliv i Masel, 1958, Nr 11, pp 66 - 68 (USSR)

**ABSTRACT:** The control of the activity of a catalyst is of importance as the activity determines its quality. The catalyst's activity can either be directly determined by cracking of a standard raw material in laboratory reactors (Refs. 11 - 13) or indirectly with the aid of their physico-chemical characteristics (Refs. 8 & 10). The described method belongs to the second group of determinations. It is based on the dependence of the catalyst's activity on its capacity of interchanging the cations contained in the same with the salt cations. The method described by Mlyussarov (Refs. 6 and 7) was applied and the exchange capacity of a number of aluminium silicate bead catalysts determined (a fresh catalyst prepared under industrial conditions; a catalyst used during catalytic cracking processes, and a catalyst which was treated with water vapour at high tem-

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ACC NR: AP6034916

gels obtained by either method are characterized by a purity of not less than  $2 \cdot 10^{-6}\%$  and may be used for the absorption of microimpurities in processing leading to a high degree of purification. [SP]

SUB CODE: 07/SUBM DATE: none/ORIG REF: 012/OTH REF: 002/

Card 2/2

ACC NR: AP6034916

SOURCE CODE: UR/0419/86/000/003/0005/0011

AUTHOR: Ogloblina, I. P. ; Krasnyy, E. B. ; Yefremov, A. A. ; Musin, T. G.

ORG: none

TITLE: Preparation and properties of high-purity silicon dioxide sorbents

SOURCE: AN BSSR. Vestsi. Seryya khimichnykh navuk, no. 3, 1966, 5-11

TOPIC TAGS: silica gel, silicon dioxide, sorbent, ethyl silicate, silicon tetrachloride, impurity

ABSTRACT: Two methods, both based on the hydrolysis of ethyl silicate and silicon tetrachloride, have been developed at the Institute of Chemical Reagents and High-Purity Substances (IREA) for obtaining high-purity synthetic silicon dioxides having a variety of adsorption properties. These are:  $\text{Si}(\text{OC}_2\text{H}_5)_4 + 4\text{H}_2\text{O} \rightarrow \text{Si}(\text{OH})_4 + 4\text{C}_2\text{H}_5\text{OH}$  and  $\text{SiCl}_4 + 4\text{H}_2\text{O} \rightarrow \text{Si}(\text{OH})_4 + 4\text{HCl}$ . The silica gels obtained from ethyl silicate showed strong water-repellent properties. A comparison of the capacity of heat treated samples to absorb benzene and water vapor showed that surface resistance to water is a function of surface dehydration. Silica

Cerd 1/2

ZHUKOVA, Z.A.; KEL'TSEV, N.V.; OGLOBLINA, I.P.; TOROCHESHNIKOV, N.S.

Using new sorbents in the advanced-stage dehydration of gases.  
Khim. ~~pr.~~ no.2:100-105 F '62. (MIRA 15:2)

(Gases--Drying)  
(Adsorbents)

Use of new absorbents for ...

S/064/62/000/002/004/008  
B101/B144

the dynamic activity of the zeolites from 21 g/100 g to 16 g/100 g. A rate of 3 l/cm<sup>2</sup>.min is assumed to be permissible for industrial adsorbers. Experiments with natural gas from the Stavropol' deposit were also conducted at 50°C and 0.5 l/min. A dew point of -60°C was reached. Because of the selective H<sub>2</sub>O vapor adsorption by zeolites the other gas components did not affect the adsorption. Even the heavy hydrocarbons do not penetrate the fine structure of the 4A zeolite pores, so that no coke formation sets in during regeneration. 200-350°C is the best regeneration temperature. For regeneration with cold gas, heating in direct flow is preferred; dew point for direct flow down to -80°C, for counter flow only -60 to -65°C. A mixture of 74.2% N<sub>2</sub>, 5.7% CO, 6.5% H<sub>2</sub>, 6.0% H<sub>2</sub>O, and 7.6% CO<sub>2</sub> was also simultaneously dried and purified. H<sub>2</sub>O was adsorbed most of all in the first zeolite layer and gradually displaced the CO<sub>2</sub> adsorbed in the following layers, which left the adsorber at a dew point below -45°C. At 12.8 g/cm<sup>3</sup> moisture, the dynamic activity of the zeolite amounted to 10 g/100 g related to CO<sub>2</sub>, and 11.3 g/100 g related to moisture. There are 8 figures, 2 tables, and 9 references: 7 Soviet and 2 non-Soviet. The reference to the English-language publication reads as follows: A. L. Kohl, F. C. Riesenfeld, Gas Purification, N. Y., 1960. ✓

Card 2/2



S/064/62/000/002/004/006  
B101/B144

AUTHORS: Zhukova, Z. A., Kel'tsev, N. V., Ogloblina, I. P.,  
Torocheshnikov, N. S.

TITLE: Use of new absorbents for intensive gas drying

PERIODICAL: Khimicheskaya promyshlennost', no. 2, 1962, 24-29

TEXT: Experiments with granulated 4A (4A) and 5A (5A) zeolites for air- and gas drying were conducted. At 20°C and 10 mm Hg, the absorptive power of these zeolite types amounted to 20.8 and 20.6 g/100 g respectively. Investigation of the adsorption isotherms of water vapor at 0-350°C showed: (1) Superiority of the zeolites compared with silica gel and aluminum oxide, (a) owing to greater moisture capacity; (b) owing to lower temperature dependence. Gases may therefore be dried by zeolites without cooling the adsorber. Experiments with an adsorber tube of 1.3 mm diameter, granulation of the zeolites 1-2 mm, depth of layer 62 cm, were conducted with air of known dew point. Results: (1) A dew point of -60 to -65°C was reached for a rate of gas flow of 1.15 l/cm<sup>2</sup>.min and 0.4 l/cm<sup>2</sup>.min in the adsorption layer. (2) Temperature increase from 30 to 80°C reduces

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Synthetic Zeolites: (Cont.)

SOV/6246

Bark, S. Ye., N. V. Kel'tsev, I. P. Ogloblina, N. M.  
Sergeyeva, M. I. Skvortsova, and N. S. Torocheshnikov.  
The Application of Synthetic Zeolites as Molecular  
Sieves for Preparing Protective Atmospheres

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AVAILABLE: Library of Congress

SUBJECT: Chemical Engineering

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BN/fmr/jk  
3/13/63

Synthetic Zeolites: (Cont.) SOV/6246

Kel'tsev, N. V., I. P. Ogloblina, and N. S. Torocheshnikov.  
 Regeneration of Zeolites in a Gas Stream 203

Vaynshteyn, S. M., G. V. Astaf'yev, Ye. Ya. Giyenko, N. I.  
 Lulova, and A. T. Slepneva. Methods of Plant and Quality  
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APPLICATION OF ZEOLITES

Kiselev, A. V., Yu. A. El'tekov, and V. N. Semenova. Ad-  
 sorption of a Mixture of Thiophene and Heptane on  
 Zeolite NaA 218

Pavlova, L. F. Adsorption From n-Hexane-Benzene Solutions  
 With Synthetic Zeolite CaA 225

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Synthetic Zeolites: (Cont.)

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COVERAGE: The book is a collection of reports presented at the First Conference on Zeolites, held in Leningrad 16 through 19 March 1961 at the Leningrad Technological Institute imeni Lensovet, and is purportedly the first monograph on this subject. The reports are grouped into 3 subject areas: 1) theoretical problems of adsorption on various types of zeolites and methods for their investigation, 2) the production of zeolites, and 3) application of zeolites. No personalities are mentioned. References follow individual articles.

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Dubinín, M. M. Introduction	5

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*С. С. ГОРБАТОВА, И. П.*

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PHASE I BOOK EXPLOITATION

SOV/6246

Soveshchaniye po tseolitam. 1st, Leningrad, 1961.

Sinteticheskiye tseolity; polucheniye, issledovaniye i primeneniye (Synthetic Zeolites: Production, Investigation, and Use). Moscow, Izd-vo AN SSSR, 1962. 286 p. (Series: Its: Doklady) Errata slip inserted. 2500 copies printed.

Sponsoring Agency: Akademiya nauk SSSR. Otdeleniye khimicheskikh nauk. Komisiya po tseolitam.

Resp. Eds.: M. M. Dubinin, Academician and V. V. Serpinskiy, Doctor of Chemical-Sciences; Ed.: Ye. G. Zhukovskaya; Tech. Ed.: S. P. Golub'.

PURPOSE: This book is intended for scientists and engineers engaged in the production of synthetic zeolites (molecular sieves), and for chemists in general.

Card 1/22 *4*

KELETNIKOV, U.V., (CHODURICH, I.P., SERGEEV, M.M., KURINOV, H.S.)

Working of gas by artificial zeolites. Study MIRA no. 35, 1978, 10  
(MIRA 14, 10)

(see 06/10)  
(see 11/10)

SHOWEN, I.S.; OGLOBLINA, I.S.

Kinetics of the carbonization of ammonium brines. Trudy  
MINTI no.35:5-11 '61. (MIRA 14:10)  
(Ammonium salts)  
(Carbonization)

OGLOBLINA, I. P., Candidate Tech Sci (diss) -- "Investigation of the kinetics of the process of carbonization of ammonia-salt solutions". Moscow, 1959. 11 pp (Min Higher Educ USSR, Moscow Order of Lenin Chem-Tech Inst im D. I. Mendelejev), 150 copies (KL, No 24, 1959, 139)



## The Carbonation Kinetics of the Ammonia Liquor

SOV/153-58-2-17/30

mixing intensity. This coefficient remains constant during the entire carbonation process at each temperature investigated and in a constant hydrodynamic treatment. The coefficient mentioned is more varying with a modification of the mixture of the solution than with temperature changes. The modified Belopol'skiy equation (2) can be applied during the whole time of carbonation for any possible temperature and way of hydrodynamic treatment. The hydrolysis of the ammonium carbamate does not form the limiting stage of the carbonation process. The rate of the hydrolysis is dependent upon the crystallization rate of the sodium bicarbonate. There are 3 figures, 2 tables, and 7 references, 5 of which are Soviet.

ASSOCIATION: Moskovskiy khimiko-tekhnologicheskii institut imeni D. I. Mendeleeva (Moscow Institute of Chemical Technology imeni D. I. Mendeleev) Kafedra tekhnologii svyazannogo azota i shchelochey (Chair of Technology of Bound Nitrogen and Alkalies)

SUBMITTED: September 9, 1957

Card 4/4

## The Carbonation Kinetics of the Ammonia Liquor

SOV/153-58-2-17/30

being always determinable, and that of the total chlorine will be sufficient. Every indication required for the calculation of  $K_a$  (constant of the absorption rate) was available. The results of this computation are presented in table 1. It can be seen from it that  $K_a$  remains sufficiently constant at a given temperature. Thus, the modified Belopol'skiy equation sufficiently determines the  $\text{CO}_2$ -absorption kinetics in the initial stage of carbonation of the ammonia liquor. The dependence of  $K_a$  upon temperature is represented by the function  $\log K_a = f\left(\frac{1}{T}\right)$  as a straight line. The activation energy calculated from the equation of this straight line is 3430 cal/mol. Since the liquid phase plays an important role in the  $\text{CO}_2$ -absorption by ammonia liquor (Ref 3), the rate of absorption will be determined in this case by the diffusion rate of free ammonia in the liquid phase towards the area separating the phases. According to the investigation of the kinetics of the 2<sup>nd</sup> carbonation stage (inclosure Fig 1) the authors come to the conclusion that the values of the absorption rate coefficient increase considerably with an increasing

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The Carbonation Kinetics of the Ammonia Liquor

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$$\sqrt{\frac{k_G}{D_{\text{liqu.}}} \cdot x_{\text{liqu.}}} \cdot k_G \text{ is the rate constant of the direct}$$

reaction, multiplied with the corresponding concentrations of compounds reacting with the absorbed substance.  $D_{\text{liqu.}}$  is the diffusion coefficient in the film of the liquid and  $x_{\text{liqu.}}$  the thickness of this film. If experimental evidence on the absorption rate of  $\text{CO}_2$  by ammonia liquor is available,  $\gamma$  can be computed according to the method of Belopol'skiy (Ref 1). The results obtained by the authors have proved that this value is close to the number 2. Thus, the value  $\beta$  is practically equal to 1 and  $\beta$  can be put equal to  $\gamma$ . The modified equation given by Belopol'skiy:  $N_A = m \cdot K_a (P_A(\text{G}) - P_A^*(\text{liqu.}))$  (2) offers no difficulties in its practical use. In order to check its applicability the authors investigated the  $\text{CO}_2$ -absorption rate by the ammonia liquor both at the beginning of the carbonation, viz. before the crystallization of the  $\text{NaHCO}_3$ , and in the final stage (at the same time with the crystallization mentioned). For the calculate of  $m$  ( $\text{NH}_4\text{OH}$ -concentration), the carbonation degree of the solution

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5(2)  
 AUTHORS: Shokin, I. N., Ogloblina, I. P. SOV/153-58-2-17/30

TITLE: The Carbonation Kinetics of the Ammonia Liquor  
 (Kinetika karbonizatsii ammiachnogo rassola)

PERIODICAL: Izvestiya vysshikh uchebnykh zavedeniy. Khimiya i khimicheskaya  
 tekhnologiya, 1958, Nr 2, pp 96-105 (USSR)

ABSTRACT: The problems under review, treated by several investigators because of their practical importance, were studied the most thoroughly by Professor Belopol'skiy (Ref 1). But the practical use of his equation (1) was complicated by the unknown quantities  $\beta$  and  $P'_A(\text{liqu.})$ . In a previous paper (Ref 7) the authors have demonstrated the ways of overcoming the difficulties regarding the latter quantity. The dimensionless parameter  $\beta$  is expressed by the following equation:

$$\beta = \frac{\sqrt{\frac{K_G}{D_{\text{liqu.}}} \cdot X_{\text{liqu.}}}}{\text{th} \sqrt{\frac{K_G}{D_{\text{liqu.}}} \cdot X_{\text{liqu.}}}} = \frac{\gamma}{\text{th} \gamma}$$

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Here,  $\gamma$  was introduced in order to cover the following term:

On the Non-Equilibrated State of the System in the Carbonization-Process of the Ammoniacal Brine (Rassol) 153-58-1-17/29

"metastable" CO<sub>2</sub> partial pressures above the solutions in the aforesaid carbonization according to a known, non-equilibrated composition of the liquid phase. The authors finally proposed a method of determination of the vapor tension above the solutions by means of a static method. There are 6 figures, 3 tables, and 9 references, 6 of which are Soviet.

ASSOCIATION: Moskovskiy khimiko-tekhnologicheskii institut imeni D. I. Mendeleyeva, Kafedra tekhnologii svyazannogo azota i shchelochey (Moscow Chemical-Technological Institute imeni D. I. Mendeleev, Chair for the Technology of Bound Nitrogen and Alkalies)

SUBMITTED: September 9, 1957

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On the Non-Equilibrated State of the System in the 153-58-1-17/29  
Carbonization-Process of the Ammoniacal Brine (Rassol)

the equilibrated solution (4), (5), (6) corresponding to it. A complete composition of the equilibrated solution can be calculated from its 3 known components for the given temperature (ref. 6). When the non-equilibrated concentrations of the 5 components: a', b', c', d' and e' are determined by analysis, complete compositions both for the non-equilibrated and the equilibrated liquid phase can be calculated at a given temperature. Results of the former for the process referred to in the title at 20, 40 and 60° are given in table 1 to 3. They are substantially different from those of the corresponding equilibrated solutions. The degree of the distance of composition of a solution from that of an equilibrated one depends on the temperature. During the whole course of the process referred to in the title, a carbamate supersaturation exists which is crystallizing in the final stage. The total vapor pressures of CO<sub>2</sub>, NH<sub>3</sub> and H<sub>2</sub>O at 20, 40, 50 and 60° above the non-equilibrated ammonia-salt solutions during their carbonization-process were determined here for the first time. An equation was deduced which allows the calculation of the

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On the Non-Equilibrated State of the System in the  
Carbonization-Process of the Ammoniacal Brine (Rassol)

153 -58-1-17/29

equilibrium of the reactions in the solution change. Taking furthermore into consideration that the crystallization of  $\text{NaHCO}_3$  involves a certain change of volume of the liquid phase, it becomes apparent that the composition of the non-equilibrated solution must differ from the equilibrated solution with respect to all components. The composition of the former can be determined rather accurately by means of analysis. The same is to a large extent also the case with ammonia-carbonate ( $\text{e}'$ ). The non-equilibrated concentrations of other components ( $\text{HCO}_3^- = \text{g}'$  non-bound ammonia =  $\text{m}'$  and the carbonic acid bound as  $\text{CO}_3^{2-} = \text{f}'$ ) can only be determined by means of calculation. For this purpose, strictly speaking, only the equations (1) and (2) can be indicated which do not connect the non-equilibrated concentrations of individual components:  $\text{d}' = \text{e}' + \text{f}' + 2\text{g}'$  (1);  $\text{m}' = \text{a}' - \text{b}' - \text{d}' + \text{g}'$  (2), in which case  $\text{a}'$  is the total quantity of  $\text{NH}_3$ ,  $\text{d}'$  = the total  $\text{CO}_2$  and  $\text{b}'$  the bound  $\text{NH}_3$ . From this, further equations for the constant ( $K_2$ ) of the reaction  $\text{HCO}_3^- + \text{OH}^- \rightleftharpoons \text{CO}_3^{2-} + \text{H}_2\text{O}$  (3) are derived for the connection between the non-equilibrated solution and

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**AUTHORS:** Shokin, I. N., Ogloblina, I. P., Solov'yeva, A. S. 153-58-1-17/29

**TITLE:** On the Non-Equilibrated State of the System in the Carbonization-Process of the Ammoniacal Brine  
(O neravnovesnom sostoyanii sistemy v protsesse karbonizatsii ammiachnogo rassola)

**PERIODICAL:** Izvestiya vysshikh uchebnykh zavedeniy. Khimiya i khimicheskaya tekhnologiya, 1958, Nr 1, pp. 108-118 (USSR)

**ABSTRACT:** In this process sodiumbicarbonate begins to precipitate from the solutions highly supersaturated with it (references 1,2). The supersaturation decreasing after the beginning of crystallization is preserved until to the end of the process of carbonization. The permanence of the non-equilibrated state in the system to be carbonized, as a whole, is caused by this. Approximating the equilibrium, not only the concentration of  $\text{HCO}_3^-$  and  $\text{Na}^+$  ions is changed in such a solution, the surplus of which is converted into the deposit, but also the concentration of other components of the solution ( $\text{CO}_3^{2-}$ ,  $\text{NH}_2\text{COO}^-$  - and  $\text{OH}^-$  -ions) since the conditions of

Card 1/4



GOLOBINA, I.P.

Using the mercury porosimeter to investigate the structure of carbon materials. Konstr. uglegraf. mat. no. 1:277-286 '64.

(NIDA 17:11)

XADULLAYEV, N.N.; SHARUTIN, A.S.; RUSTAMBEKOV, A.F.; OGLOBLINA, G.P.;  
ALIYEV, T.A.; VANYAVKIN, B.P.; GURVICH, M.L.

Oil well drilling in the Kyursangya area. Bureau no. 3:7-10 '65.  
(MIRA 18:5)

1. "AzNIItburneft" i brest "Aznefterazvedka".

L

OGLOBLIN, Ye. S., Cand Agr. Sci -- (diss) "Certain problems  
of the economics and organization of sheep <sup>breeding</sup> ~~breeding~~ under  
conditions of extensive <sup>plow</sup> ~~plows~~ <sup>of lands</sup> ~~of soils fit for cultivation~~  
(Based on the example of the sheep <sup>sovkhozes</sup> ~~sovkhozes~~ (sovkhozes)  
of Northern Caucasus)." Mos, 1957. 18 pp including cover.  
(All-Union Sci Res Inst of Animal Husbandry. Department of  
Economics and Organization of Animal Husbandry.) 110 copies.  
(KL, 12-58, 100)

*sheep farms*

CATEGORY : Forestry, Dendrology.

ABS. JOUR: *Ref Zhur-Biologiya*, No. 5, 1959, No. 20129

AUTHOR : Ogloblin, Ya.S.

INST. : Inst. of Forestry, Acad. of Sciences USSR

TITLE : Wartybark Euonymus on the Steep Chalk Slopes  
of the Mid Desna River Basin.

ORIG. PUB.: *Tr. In-ta lesa. AN SSSR*, 1958, 46, 25-34

ABSTRACT : Young deciduous scrub woods on compost-carbonate soils of the steep chalk slopes of the middle Desna River basin, made up of ash, Norway maple, linden, alder, filbert, Tatarian honeysuckle, and other trees, are typified by abundant participation of wartybark euonymus (up to 40% of the total amount of scrub species) in the underwood. The latter thrives better here than in other environments, its shrubs, displaying great individual variations in girth

REF. ID: A66112

OGLEBIN, Ya.S.

New method for determining the gutta content of spindle tree bark  
and possible industrial utilization of bark waste. Trudy Inst.  
lesa 46:20-24 '58. (MIRA 11:6)

1. Bryanskiy lesokhozyaystvennyy institut.  
(Gutta-percha) (Spindle tree) (Fibers)

USSR / Forestry. Forest Management

K-4

Abs Jour: Ref Zhur-Biol., No 10, 1958, 43943

Author : Ogloblin, Ya. S.

Inst : Bryansk Forest Management Institute

Title : Construction of a Device for the Measurement of  
the Crown Projections

Orig Pub: Tr. Bryanskogo lesokhoz. in-ta, 1957, 8, 257-259

Abstract: This article describes the layout and the method of using an uncomplicated device constructed for precise measurement of the radius of the crown rising considerably above the ground. It is noted that for the height of 10 meters the margin of error does not exceed  $\pm 3-5$  centimeters. A schematic drawing of the device is shown.

Card 1/1

17

USSR /Chemical Technology. Chemical Products  
and Their Application

-27

Wood chemistry products. Cellulose and its  
manufacture. Paper.

Abs Jour: Referat Zhur - Khimiya, No 9, 1957, 32695

cellulose suitable for the manufacture of filter  
paper, alcohol, etc.

Card 2/2

*Ogloblin, Ya. S.*

USSR /Chemical Technology. Chemical Products  
and Their Application

I-27

Wood chemistry products. Cellulose and its  
manufacture. Paper.

Abs Jour: Referat Zhur - Khimiya, No 9, 1957, 32695

Author : Ogloblin Ya, S.

Inst : Bryansk Institute of Forestry

Title : Method for the Determination of Gums and  
Cellulose in Small Samples

Orig Pub: Tr. Bryanskogo leskhoz. in-ta, 1956, 7, 202-204

Abstract: Description of a method for simultaneous deter-  
mination of gums and cellulose in the bark of  
euonymus. One ton of bark yield about 500 kg

Card 1/2



KASHCHAYEV, V.V., inzh.; ~~UGLOVIN, V.N., inzh.~~

Reducing accidents in Kirghizia mines. Bezop. truda v prom. 5  
no. 2:14-15 F '61. (MIRA 14:2)

1. Gosgortekhnadzor Kirgizskoy SSR.  
(Kirghizistan--Mining engineering--Safety measures)

OGLUBLIN, S.I.

Automatizing the process of knocking-out casting boxes, and shifting the casting grills from pneumatic to electric drive. Prom.energ. 10 no.5:4-6  
My '53. (MLRA 6:5)

(Founding)

OGIOBLIN, S.A., kand. tekhn. nauk; OL'BINSKIY, Z.M., inzh.,  
retsenzent; ZHURAVLEV, S.A., kand. tekhn. nauk, red.

[Dividing heads and their adjustment] Delitel'nye golovki  
i ikh nastroyki. Pod red. S.A. Zhuravleva. Moskva,  
Mashinostroenie, 1964. 136 p. (Bibliotechka frezerovshchika,  
no.6) (MIRA 19:1)

VOLKOV, G.K., assistént; SHESTERKINA, N.V.; OGLOBLIN, N.Ye.; KASYUK, I.I.,  
veterinarnyy vrach

Air ionization increases the sexual activity of bulls. Veterinariia  
40 no.9:47-48 S '63. (MIRA 17:1)

1. Moskovskaya veterinarnaya akademiya (for Volkov). 2. Zaveduyushchaya  
Stupinsky stantsiyey iskusstvennogo osemneniya sel'skokhozyaystvennykh  
zhivotnykh, Moskovskoy oblasti. (for Shesterkina). 3. Starshiy zootekh-  
nik-laborant Stupinskoy stantsii iskusstvennogo osemneniya sel'ske-  
khozyaystvennykh zhivotnykh, Moskovskoy oblasti (for Ogloblin).

BELOV, K.A.; VOLKOVA, O.B.; MAKSIMOVA, M.I.; OGLOBLIN, N.D.; LUK'YACHENKO,  
V.N.; TUL'CHINSKAYA, A.Ya.

Effect of the chemical composition of the reagents, used for coal  
flotation, on their activity. Koks i khim. no.8:8-12 '62.  
(MIRA 17:2)

1. Khar'kovskiy politekhnicheskiy institut (for Belov, Volkova,  
Maksimova). 2. Khar'kovskiy gornyy institut (for Ogloblin, Luk'-  
yanchenko, Tul'chinskaya).

OGLOBLIN, N.D., inzh.

New reagent for the flotation of Donets coals. Izv. vys. ucheb.  
zav.; gor. zhur. no.11:167-172 '61. (MIRA 15:1)

1. Khar'kovskiy gornyy institut. Rekomendovana kafedroy obogash-  
cheniya poleznykh iskopayemykh.  
(Donets Basin--Coal) (Flotation--Equipment and supplies)

OGLOBLIN, N.D.

Reagent for the flotation of coal smalls. *Biol.tekh.-ekon.inform.*  
no.7:17-19 '61. (MIRA 14:8)  
(Coal preparation)

OGLOBLIN, N.D.; LUK'YANCHENKO, V.N.

Testing of a new reagent for the flotation of slurries. Koks i  
khim. no. 3418-21 '61. (MIRA 14:4)

1. Khar'kovskiy gornyy institut,  
(Flotation) (Coal preparation)



OGLOBLIN, N.D., inzh.

Behavior of pyrite in the flotation of coal pulp. *Izv. vys. ucheb. zav.; gor. zhur. no. 6: 132-138 '60.* (MIRA 14:5)

1. Khar'kovskiy gornyy institut. Rekomendovana kafedroy obogasheniya poleznykh iskopayemykh. (Coal preparation) (Flotation)

OGLOBLIN, N.D., inzh.

U.S. GOVERNMENT PRINTING OFFICE: 1963 O 282 000

Use of the Shebelinka gas condensate as a reagent in the  
flotation of fine coals. Ugol' Ukr. 4 no. 11:15-17 N '60.

(MIRA 13:12)

(Flotation--Equipment and supplies)

11(7)

PHASE I BOOK EXPLOITATION

SOV/2174

Oglobin, Nikolay Dmitriyevich, and Igor' Viktorovich Trushlevich

Tekhnicheskii kontrol' na ugleobogatitel'nykh fabriках (Technical Control in Coal Preparation Plants) Moscow, Ugletekhizdat, 1959. 210 p. Errata slip inserted. 4,000 copies printed.

Resp. Ed.: K. A. Korovenkova.

**PURPOSE:** This textbook is intended for students in mining schools and institutes. It may also serve as a manual for the technical control of personnel of coal preparation plants.

**COVERAGE:** This textbook describes the operation of a modern coal preparation plant with its mechanical equipment and procedures used to improve the quality of run-of-mine coal. In addition to general information on coal preparation, the authors describe in detail coal sampling, crushing, screening, dedusting, dewatering, drying, weighing, etc. They set forth the scientific

Card ~~1/7~~

OGLOBLIN, L.

Grain transshipping practices in the port of Leningrad.  
Mor. flot 24 no.2:8-11 F '64.

(MIRA 18:12)

1. Pomoshchnik nachal'nika Leningradskogo porta po  
mekhanizatsii.

KONDRAT'YEV, K.P.; OGLOBLIN, L.A.; SKOMOROVSKIY, R.V., spetsred.;  
DENISOV, K.N., red. izd-va; DROZHEVINA, L.P., tekhn.red.

[Operating hydraulic systems of harbor transloading machines]  
Eksploatatsiia gidravlicheskikh sistem portovykh peregruzochnykh  
masin. Leningrad, Izd-vo "Morskoi transport," 1960. 127 p.  
(MIRA 14:4)

(Cargo handling)

(Oil hydraulic machinery)

OGLOBLIN, L. A.

118-58-3-5/21

**AUTHORS:** Ogloblin, L.A., and Belokopytov, V.A., Engineers

**TITLE:** The MVS-1 Railroad Car Unloading Machine for Loose Goods  
(Vagonorazgruzochnaya mashina dlya slezhivayushchikhsya  
sypuchikh gruzov MVS-1)

**PERIODICAL:** Mekhanizatsiya Trudoyemkikh i Tyazhelykh Rabot, 1958, # 3,  
pp 16-17 (USSR)

**ABSTRACT:** Plants of the Ministry of Merchant Marine have constructed a railroad-freight-car unloading machine of the type MVS-1, designed by the TsPKB-4 of the Ministry and tested at the port of Osipenko. The machine mechanizes the unloading operation of freight, such as salt, superphosphate, etc, and considerably speeds up the unloading work, especially when combined with conveyors or grab cranes. Parameters of this machine are given as follows: working capacity - 65 tons per hour; dimensions - 4,850x1,730x1,965 mm; weight - 2,425 kg; conveyor extension - 2,500 mm and the width of the belt - 400 mm.

The unloading of a 60-ton freight car of salt by manual work requires 4 laborers and lasts from 4 to 5 hours. The use of the new machine cuts unloading time by "2 to 2.5

Card 1/2

OGLOBLIN, L.

Present state of the means of mechanization and its maintenance by integrated brigades. Mor. flot 23 no.8:13-14  
Ag '63. (MIRA 16:11)

1. Pomoshchnik nachal'nika Leningradskogo porta po mekhanizatsii.

OGLOBLIN, K.A.; SEMENOV, V.P.

Reaction of nitrosyl chloride with unsaturated compounds.  
Part 23: Reaction involving 3,3,1-trichloro-2-methylpropene  
and 2-nitropropene. Zhur. org. khim. J. no.8:1356-1360 1965.

Reaction of nitrosyl chloride with unsaturated compounds.  
Part 24: Reaction involving methyl ester and nitrile of  
acrylic acid. Ibid.:1361-1364 (NIR: 13:11)

1. Leningradskiy gosudarstvennyy universitet.



CHIGELIN, K.A.; POTEKIN, A.A.

Reaction of nitrosyl chloride with unsaturated compounds.  
Part 12: Reaction involving  $\alpha, \beta$ -unsaturated aldehydes.  
Zhur. org. khim. 1 no.8:1352-1356 Ag '65. (USSR 18:11)

L. I. Lenin skly gosudarstvennyy universitet.

SAMARTSEV, M.A.; OGLOBLIN, K.A.

Interaction of nitrosyl chloride with unsaturated compounds. Part 17:  
Cis-trans isomerism of dimers of aliphatic chloronitroso compounds.  
Zhur. org. khim. 1 no.1:31-37 Ja '65. (MIRA 18:5)

1. Leningradskiy gosudarstvennyy universitet.

OGLOBLIN, K.A.; SEMENOV, V.P.

Interaction of nitrosyl chloride with unsaturated compounds.  
Part 16: Reaction with methyl methacrylate. Zhur. org. khim.  
1 no.1:27-30 Ja '65. (MIRA 18:5)

1. Leningradskiy gosudarstvennyy universitet.

OCLOBLIN, K.A.; SAMARTSEV, M.A.

Interaction of nitrosyl chloride with unsaturated hydrocarbons.  
Part 2: Reaction with 2-pentene. Zhur. ob. khim. 34, no. 5,  
1525-1530 My '64.

Interaction of nitrosyl chloride with unstaunated hydrocarbons.  
Part 13: Reaction with 2,2-dimethyl-3-pentene. Ibid. s1530-1536  
(MIRA 17:77)

1. Leningradskiy gosudarstvennyy universitet.

OGLOBLIN, K.A.; SEMENOV, V.P.

Interaction of nitrosyl chloride with unsaturated hydrocarbons.  
Part 11: Reaction with 2-chloropropene. Zhur. ob. khim.  
no. 5:1522-1525 My '64. (MIRA 1967)

1. Leningradskiy gosudarstvennyy universitet.

OGLOBLIN, K.A.; POTEKHIN, A.A.

Synthesis and cleavage of dimers of  $\beta$ -chloro- $\alpha$ -nitroso  
carbonyl compounds. Dokl. AN SSSR 159 no.4:853-856 D '64  
(MIRA 18:1)

1. Leningradskiy gosudarstvennyy universitet im. A.A. Zhdanova.  
Predstavleno akademikom A.N. Nesmeyanovym.

OGLOBLIN, K.A.; SEMENOV, V.P.

Reaction of vinyl ethers with nitrosyl chloride. Dokl. AN SSSR  
155 no. 1 145-148 Mr '64. (MIRA 17:4)

1. Leningradskiy gosudarstvennyy universitet im. A.A.Zhdanova.  
Predstavleno akademikom A.N.Nesmeyanovym.

SAMARTSEV, M.A.; OGLOBLIN, K.A.

Preparation of cis- and trans-isomers of dimeric chloronitroso  
compounds. Zhur.ob.khim. 34 no.1:361 Ja '64. (MIRA 17:3)

1. Leningradskiy gosudarstvennyy universitet.



OGLOBLIN, K.A.; KALIKHEVICH, V.N.; POTEKHIN, A.A.; SEMENOV, V.P.

Interaction of nitrosyl chloride with unsaturated hydrocarbons. Part  
9: Reaction with mono- and assym. disubstituted ethylenes. Zhur.ob.  
khim. 34 no.1:170-181 Ja '64. (MIRA 17:3)

1. Leningradskiy gosudarstvennyy universitet.

OGLOBLIN, K.A.; SAMARTSEV, M.A.

Reaction of nitrosyl chloride with unsaturated hydrocarbons.  
Part 8: Reaction with 2-butene. Zhur.ob.khim. 33 no.10:3257-  
3262 0 '63. (MIRA 16:11)

1. Leningradskiy gosudarstvennyy universitet.

OGLOBLIN, K.A.; SEMENOV, V.P.; SKOBLIKOVA, V.I.

Interaction of nitrosyl chloride with unsaturated hydrocarbons.  
Part 7: Conversion of nitrosyl chlorides of olefins to  
oximes of  $\alpha$ -chloro aldehydes and  $\alpha$ -chloro ketones brought  
about by the action of hydrogen chloride. Zhur.ob.khim.  
33 no.3:888-896 Mr '63. (MIRA 16:3)

1. Leningradskiy gosudarstvennyy universitet.  
(Nitrosyl chloride)  
(Olefins) (Oximes)

OGLOBLIN, K.A.; FOTEKHIN, A.A.

Interaction of nitrosyl chloride with diallyl and bivinyl.  
Zhur.ob.khim. 31 no.7:2438-2439 J1 '61. (MIRA 14:7)

1. Leningradskiy gosudarstvennyy universitet imeni A.A. Zhdanova.  
(Nitrosyl chloride) (Hexadiene) (Butadiene)

Reaction of Isobutylene Nitrosochloride  
With Amines. Synthesis of  $\alpha$ -Amino-  
aldehydes

78263  
SOV/79-30-3-17/69

Key to Table 1. (1) Nr of compound; (2) mp; (3) found;  
(4) calculated; (5) neutral equivalent; (6) yield (in %),  
based on nitrosochloride.

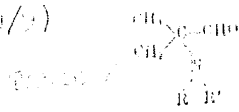
Key to Table 2. (1)  $\alpha$ -Aminoaldehydes; (2) Nr of  
compound; (3) bp (pressure in mm); (4) found; (5) cal-  
culated; (6) neutral equivalent; (7) yield (in %);  
(8) aldehyde; (9) dimer.

ASSOCIATION: Leningrad State University (Leningradskiy gosudarstvennyy  
universitet)

SUBMITTED: March 31, 1959

Card 9/9

(Key to Table 2 on Card 9/3)



78263 30V/19-50-3-17/00

2	R	R'	3	%	d <sub>r</sub>	MP <sub>r</sub>		6		7	
						4	5	4	5	8	9
(X)	CH <sub>3</sub>	H	—	—	—	—	—	—	—	—	60
(XI)	C <sub>2</sub> H <sub>5</sub>	H	40 (29)	1.4532	0.9398	33.16	33.44	109.3	115.1	8	18
(XII)	n-C <sub>3</sub> H <sub>7</sub>	H	62-62.5 (23)	1.4561	—	—	—	129.1, 130.5	129.1	2	57
(XIII)	n-C <sub>4</sub> H <sub>9</sub>	H	58-62 (5)	1.4583	0.9282	42.12	42.72	140.9, 142.8	143.2	15	32
(XIV)	CH <sub>3</sub> C <sub>2</sub> H <sub>5</sub>	H	65 (18)	1.4309	0.8659	42.81	42.56	146.1, 147.0	143.2	68	—
(XV)	CH <sub>3</sub> C <sub>2</sub> H <sub>5</sub>	H	70 (9)	1.4364	0.8601	52.08	51.99	171.1, 172.0	171.2	33	—
(XVI)	CH <sub>3</sub>	CH <sub>3</sub>	57-58 (54)	1.4290	0.8817	33.97	33.58	118.5, 119.2	115.1	51	—
(XVII)	C <sub>2</sub> H <sub>5</sub>	C <sub>2</sub> H <sub>5</sub>	59.5 (18)	1.4311	0.8653	42.85	42.87	143.7, 144.2	143.2	67	—

8/3

(Ref. to Table I on sheet 2, 9)

Case 101/10 10-9-11

Table I. Reaction products of isobutylene nitrono-chloride with anion.

i	$\begin{matrix} \text{Cl} \\   \\ \text{CH}_2 - \text{C} - \text{CH}_2 - \text{N} \\   \quad   \\ \text{R} \quad \text{R}' \end{matrix}$		2	δ		δ'		6
	R	R'		3	4	3	4	
(I)	CH <sub>3</sub>	H	98-99 <sup>a</sup>	24.35	24.12	115.1, 116.6	115.1	81
(II)	C <sub>2</sub> H <sub>5</sub>	H	108-108.5	21.62	21.53	129.5, 130.4	130.1	68
(III)	n-C <sub>3</sub> H <sub>7</sub>	H	78.5-79.5	19.39, 19.57	19.43	143.1, 143.2	144.2	82
(IV)	n-C <sub>4</sub> H <sub>9</sub>	H	70-71	...	...	158, 158.3	158.2	67
(V)	$\begin{matrix} \text{CH}_3 \\   \\ \text{CH} \\   \\ \text{C}_2\text{H}_5 \end{matrix}$	H	48-49	17.38, 17.48	17.71	158.2, 158.4	158.2	59
(VI)	$\begin{matrix} \text{CH}_3 \\   \\ \text{CH} \\   \\ \text{C}_3\text{H}_7 \end{matrix}$	H	51-52	14.87	15.04	185.2, 185.3	184.3	59
(VII)	CH <sub>3</sub>	H	126-127	14.36, 14.69	14.58	192.0, 192.6	192.3	43
(VIII)	CH <sub>3</sub>	CH <sub>3</sub>	95-95.5	21.33, 21.64	21.53	130.4, 131.1	131.1	38
(IX)	C <sub>2</sub> H <sub>5</sub>	C <sub>2</sub> H <sub>5</sub>	68-69	17.65, 17.72	17.71	156.1, 159.2	158.2	70

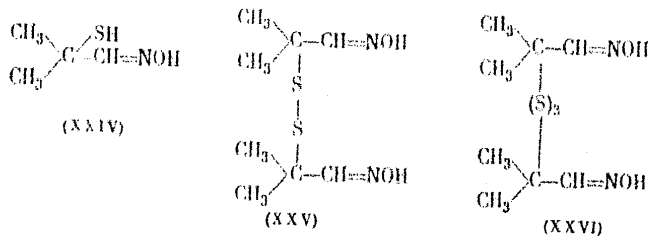
Card 7/9

Reaction of Isobutylene Nitrosochloride  
 With Amines. Synthesis of  $\alpha$ -Amino-  
 aldehydes

78263

SOV/79-30-3-17/69

The reaction of isobutylene nitrosochloride with NaHS yields compounds (XXV) (yield 10%), mp 131.5-132<sup>o</sup>; and (XXVI) (yield 16%), mp 149.5-150.5<sup>o</sup>, instead of the expected (XXIV).



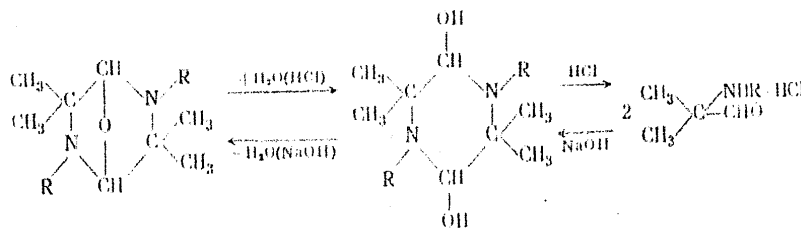
There are 1 figure; 2 tables; and 24 references, 7 Soviet, 5 U.S., 3 U.K., 7 German, 2 French. The 5 U.S. references are: Balenovic, K., Bregant, N., Cerar, D., Fles, D., J. Org. Chem., 18, 297 (1953); U.S. Patent 2494859; Mellor, J., Inorg. & Theor. Chem., 8, 289 (1947); Inorganic Synthesis, N.Y., 4, 48 (1953); Iffland, C., Teh-Fu-Uen, J. Am. Chem. Soc., 76, 4180 (1954).

Card 6/9

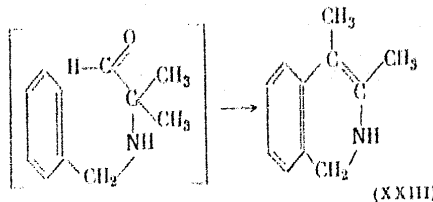


Reaction of Isobutylene Nitrochloride  
 With Amines. Synthesis of  $\alpha$ -Amino-  
 aldehydes

78263  
 SOV/19-30-3-17/6

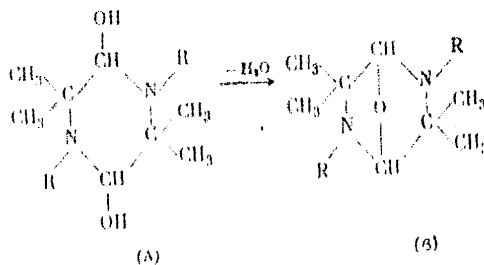


Heating the hydrolysis product of (VII) with concentrated H<sub>2</sub>SO<sub>4</sub> at 180-220<sup>o</sup> yields a small amount of compound (XXIII), mp of its picrate, 212-213<sup>o</sup>.



Reaction of Isobutylene Nitrosochloride  
With Amines. Synthesis of  $\alpha$ -Amino-  
aldehydes

78263  
SOV/79-30-3-17/69

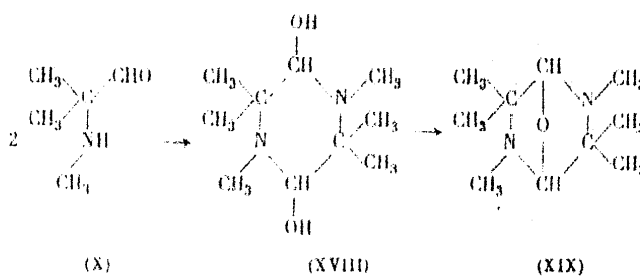


Dimers (XIX)-(XXII) are stable in organic solvents, but in an aqueous acidified medium they are present, at least partially, in the form of monomolecular aminoaldehydes. The reaction of dimer (XXI) with an alcoholic solution of  $\text{NH}_2\text{OH} \cdot \text{HCl}$  yields the hydrochloride of aminoaldoxime (III). The majority of dimers reduce both Tollen's and Fehling's agent. On the basis of this, it is assumed that there is an equilibrium between the monomer and dimer forms of  $\alpha$ -aminoaldehydes.

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Reaction of Isobutylene Nitrosochloride  
 With Amines. Synthesis of  $\alpha$ -Amino-  
 aldehydes

78263  
 SOV/79-30-3-17/69

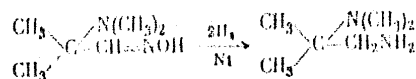


Aminoaldehydes (XI), (XII), and (XIII) were obtained in monomer form in low yields. The basic products in this case were oxides (XX), bp 42-43<sup>o</sup>; (XXI), mp 178-178.5<sup>o</sup>; and (XXII) formed as a result of removal of water from compound (A).

Reaction of Isobutylene Nitrosochloride  
 With Amines. Synthesis of  $\alpha$ -Amino-  
 aldehydes

78263  
 SOV/79-30-3-17/69

amino group in the reaction products of primary amines was confirmed by formation of nitrosoamines in reaction of  $\alpha$ -aminoaldehydes with nitrous acid. The obtained oximes can be reduced to diamines over skeletal nickel in alcohol:



To avoid formation of oximes during the alkalization of the hydrolysis products of aminoaldehydes, sodium nitrite was introduced into the solution to decompose hydroxylamine. Aminoaldehydes were extracted with ether. The yield and physical constants of the obtained aldehydes are shown in Table 2. Compound (X) was not separated, but two corresponding dimers were obtained ((XVIII), mp 59.5-61° and (XIX), bp 77.5-79.5° (13 mm),  $n_D^{20}$  1.4571).

Card 2/9

5.3610

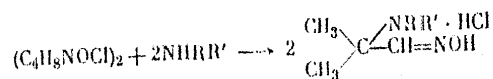
78203  
30V/79-30-3-17/69

AUTHORS: Ogloblin, K. A., Samartsev, M. A

TITLE: Reaction of Isobutylene Nitrosochloride With Amines.  
Synthesis of  $\alpha$ -Aminoaldehydes

PERIODICAL: Zhurnal obshchey khimii, 1960, Vol 30, Nr 3,  
pp 805-820 (USSR)

ABSTRACT: Isobutylene nitrosochloride was reacted with methyl-ethyl-, n-propyl-, n-butylamine, 2-aminobutane, 2-aminohexane, benzylamine, dimethyl- and diethylamine.  $\alpha$ -Aminoaldoximes and their salts were the main reaction products. The general scheme of the reaction is:



The physical constants of the obtained  $\alpha$ -aminoaldoximes are shown in Table 1. Aqueous and alcoholic solutions of  $\alpha$ -aminoaldoximes produce a color reaction with salts of heavy metals. The presence of the secondary

Card 1/3

Reaction of Nitrosyl Chloride With Unsaturated  
Hydrocarbons. VI. Reaction With  $\alpha$ -Methyl Styrene

SOV/79-29-6-51/72

ASSOCIATION: Leningradskiy gosudarstvennyy universitet (Leningrad  
State University)

SUBMITTED: May 12, 1958

Card 3/3

Reaction of Nitrosyl Chloride With Unsaturated  
Hydrocarbons. VI. Reaction With  $\alpha$ -Methyl Styrene

SOV/79-29-6-51/72

derivatives (I), (II), (III) are the main products, which points to the predominant effect of nitrosyl chloride on the  $\alpha$ -methyl styrene. Apparently (I) results due to the cleavage of hydrochlorides of (III) and (II) in the chlorination of the  $\alpha$ -methyl styrene and according to L'vov-Tishchenko (Ref 2 - Scheme 1). From the monochlorides which cannot be separated by distillation, HCl, formaldehyde, formic acid, acetophenone, and chloro acetophenone were obtained by ozonolysis (Scheme 2). The yield of compound (II) was 43 %, computed for the amounts of formaldehyde and formic acid obtained. The saponification of the monochloride fraction with alcoholic alkali lye yielded similar results. The formation of compound (IV) may be explained by scheme 3, i. e. by oxidation of the primary reaction product (V) with NOCl to propane (VI). The presence of acetophenone in the reaction products is explained by the hydrolysis of the compound in the distillation of (IV) (Scheme 4). There are 18 references, 6 of which are Soviet.

Card 2/3

5 (3)

## AUTHORS:

Ogloblin, K. A., Semenov, V. P.

SOV/79-29-6-51/72

## TITLE:

Reaction of Nitrosyl Chloride With Unsaturated Hydrocarbons  
(Vzaimodeystviye khloristogo nitrozila s nenasyshchennymi  
uglevodorodami). VI. Reaction With  $\alpha$ -Methyl Styrene  
(Reaktsiya s  $\alpha$ -metilstirolom)

## PERIODICAL:

Zhurnal obshchey khimii, 1959, Vol 29, Nr 6,  
pp 2006-2012 (USSR)

## ABSTRACT:

The problem of the reaction of nitrosyl chloride with mono- and unsymmetric, double-substituted ethylenes has not been sufficiently dealt with in publications. In order to investigate the influence exercised by the substituents at the ethylene bond on the direction of the reaction, a reaction of nitrosyl chloride with  $\alpha$ -methyl styrene was carried out. As was done already earlier (Ref 1), the reaction was carried out in ether solution with equimolecular reagents or with a small nitrosyl chloride excess. The following compounds were separated from the reaction products: the mixture of two unsaturated monochlorides (I) and (II) which cannot be separated by distillation, the compounds (III) and (IV). In some low fractions acetophenone was found. Mono- and dihalogen

Card 1/3



5(3)  
AUTHOR: Ogloblin, K. A. SOV/79-29-5-74/75

TITLE: The Preparation of Acid Chlorides of  $\alpha$ -Chloroisobutyric and  $\alpha$ -Chloro- $\alpha$ -Methylbutyric Hydroxamic Acids From 2-Methylpropene and 2-Methylbutene-1 (Polucheniye khlorangidridov  $\alpha$ -khlorizobutiro- i  $\alpha$ -khloro- $\alpha$ -metilbutirogidroksamovykh kislot iz 2-metilpropena i 2-metilbutena-1)

PERIODICAL: Zhurnal obshchey khimii, 1959, Vol 29, Nr 5, pp 1752-1753 (USSR)

ABSTRACT: The compounds mentioned were produced in ether solution in the presence of hydrogen chloride. Physical data are specified (Table). Furthermore, the methyl esters of these acids were prepared. On the preparation of  $\alpha$ -chloroisobutyrohydroxamic acid also 3-chloro-2-methylpropene-1 and 1,2-dichloro-2-methylpropane were formed in small quantities. There is 1 table.

ASSOCIATION: Leningradskiy gosudarstvennyy universitet (Leningrad State University)

SUBMITTED: March 17, 1959  
Card 1/1

Reaction of Nitrosyl Chloride With Unsaturated Hydro- SOY/79-29-1-56/74  
 carbons. V. Reaction of Butene-1 and 2-Methyl Butene-3 With Nitrosyl Chloride  
 in the Presence of Hydrogen Chloride. Preparation of Chloric Acid Anhydride  
 of the  $\alpha$ -Chloro Isovalero- and  $\alpha$ -Chloro Butyro Hydroxamic Acid

hydroxamic acid. This preparation is new and is apparently to be regarded as a general method for the synthesis of chloric acid anhydrides of  $\alpha$ -chloro hydroxamic acid from mono-substituted ethylenes. Beside these, inconsiderable quantities of saturated dichlorides are formed as e.g. 3,4-dichloro-2-methyl butene and others. As it can be seen hydrogen chloride plays an important part in this reaction. Under its influence the regrouping of the initial products, the chloro-nitroso compounds into oximes of  $\alpha$ -chloro aldehydes takes place which in turn yield together with nitrosyl chloride the chloric acid anhydrides of  $\alpha$ -chloro hydroxamic acids. There are 14 references, 7 of which are Soviet.

ASSOCIATION:

Khimicheskii institut Leningradskogo gosudarstvennogo universiteta (Chemical Institute of the Leningrad State University)

SUBMITTED:  
 Card 3/3

November 20, 1957

Reaction of Nitrosyl Chloride With Unsaturated Hydro- SOV/79-29-1-56/74  
carbons. V. Reaction of Butene-1 and 2-Methyl Butene-3 With Nitrosyl Chloride  
in the Presence of Hydrogen Chloride. Preparation of Chloric Acid Anhydride  
of the  $\alpha$ -Chloro Isovalero- and  $\alpha$ -Chloro Butyro Hydroxamic Acid

chloride with 2-methyl butene-3 (Ref 2) and butene-1 (Ref 3) no solid nitroso chloride is formed. The main products are the corresponding 1-nitro-2-chloro alkanes i.e. the initially forming monomer nitroso chloride is oxidized by nitrosyl chloride to the chloro-nitro compound, whereas the chlorine forming in connection with the decomposition of nitrosyl chloride is added to the hydrocarbon and forms dichloride. Consequently, no regrouping of the nitroso-chlorides of the mono-substituted ethylene into chlorooximes takes place as their oxidation into the nitro compounds proceeds more rapidly. In order to be able to continue the investigations of the influence of HCl upon the regrouping of the chloro-nitroso compound into chlorooxime the reaction of nitrosyl chloride with 2-methyl butene-3 (I) and butene-1 (II) was carried out in the presence of hydrogen chloride. The result was 50% chloric acid anhydride of  $\alpha$ -chloro isovalero hydroxamic acid and  $\alpha$ -chloro butyro

Card 2/3

AUTHOR:

Ogloblin, K. A.

SOV/79-29-1-56/74

TITLE:

Reaction of Nitrosyl Chloride With Unsaturated Hydrocarbons (Vzaimodeystviye khlorigo nitrozila s nenasyshchennymi uglevodorodami)V. Reaction of Butene-1 and 2-Methyl Butene-3 With Nitrosyl Chloride in the Presence of Hydrogen Chloride. Preparation of Chloric Acid Anhydride of the  $\alpha$ -Chloro Isovalero- and  $\alpha$ -Chloro Butyro Hydroxamic Acid (Reaktsiya butena-1 i 2-metilbutena-3 s khlorigim nitrozilom v prisutstvii khlorigo vodoroda. Polucheniye khlorangidridov  $\alpha$ -khlorigovalero- i  $\alpha$ -khlorbitirogidroksamovoy kislot)

PERIODICAL:

Zhurnal obshchey khimii, 1959, Vol 29, Nr 1,  
pp 264 - 269 (USSR)

ABSTRACT:

In one of the previous reports it was shown (Ref 1) that in the case of reaction of nitrosyl chloride with trimethyl ethylene in the presence of HCl not the dimer of trimethyl ethylene nitroso-chloride is formed as the main product but the monomer oxime of 2-chloro-2-methyl butanone-3. The presence of HCl favors the regrouping of the initially forming monomer chloro-nitroso compounds into chlorooxime (Scheme 1). In connection with the reaction of nitrosyl

Card 1/3

OGIOBLIN, K.A.; MARKINA, G.V.

Chromatographic separation of mixtures of nitro- and dinitro-  
benzoic acids. Vest. LGU 14 no. 4:149-153 '59. (MIRA 12:5)  
(Benzoic acid) (Chromatographic analysis)

SOV/54-59-1-22/25

## Chromatographic Separation of Mixtures of Nitro- and Dinitrobenzoic Acids

acids is based upon their arrangement in the column in the following order: o- > m-, o- > p-, m- > p-, m- > 2,4-, 3,5- > m-, p- > 2,4-, 3,5- > 2,4-. This arrangement corresponds to the dependence of the adsorbability on the intensity of the acid, which was detected already earlier. The dissociation constants of the various acids are given in a table. The other three mixtures investigated were anomalous compared with the dependence detected here. The individual separation processes are described in detail in the experimental part. There are 1 table and 13 references, 7 of which are Soviet.

SUBMITTED: May 30, 1958

Card 2/2

AUTHORS: Ogloblin, K. A., Markina, G. V. SOV/54-59-1-22/25

TITLE: Chromatographic Separation of Mixtures of Nitro- and Dinitrobenzoic Acids (Khromatograficheskoye razdeleniye smesey nitro- i dinitrobenzovnykh kislot)

PERIODICAL: Vestnik Leningradskogo universiteta. Seriya fiziki i khimiya, 1959, Nr 1, pp 149-153 (USSR)

ABSTRACT: The chromatographic analysis used here is connected with earlier investigations concerning the chromatographic separation of mixtures of other nitroderivatives of the aromatic series (Refs 1-4). An aqueous oxide of aluminum with a water content of 5 - 17% was used for the separation of the mixture mentioned in the title, and solvents were benzene and mixtures of benzene with ether and acetone. Experiments with other solvents did not yield satisfactory results. The investigated mixture contained o-(I), m-(II), and n-(III)-nitrobenzoic acids, 2,4-(IV) and 3,5-(V)-dinitrobenzoic acids. The separation was successfully carried out in the double mixtures (I)+(III), (II)+(III), (II)+(IV), (III)+(IV), (IV)+(V). From the mixtures (I)+(II) and (II)+(V) only (II) was obtained in pure form. It is less adsorbable than (I) and (V). The adsorbability of the individual

Card 1/2

SOV/79-28-12-18/41

Reaction of Nitrosyl Chloride With Unsaturated Hydrocarbons. IV. Reaction With  
Butene-1

oxidizes with nitrosyl chloride to the 1-nitro-2-chloro butane. The chlorine formed in the decomposition of the nitrosyl chloride yields with hydrocarbon the dichloride (Scheme 1). The absence of the oxime of  $\alpha$ -chloro butyric acid aldehyde in the reaction products indicates a low velocity of regrouping of the nitroso compound, as compared to that of the oxidation process. The presence of the acid chloride of  $\alpha$ -chloro-butyro-hydroxamic acid in the reaction products indicates the formation of chlorine oxime. This acid chloride apparently is formed on the action of chlorine and nitrosyl chloride on the oxime (Ref 3) (Scheme 2). The structure of 1-nitro-2-chloro butane was proved by its transformation into  $\alpha$ -chloro butyric acid on heating it in a sealed ampoule at 140-145<sup>o</sup> (Ref 1) (Scheme 3).- There are 9 references, 2 of which are Soviet.

ASSOCIATION: Khimicheskiy institut Leningradskogo gosudarstvennogo universiteta  
(Chemical Institute of the Leningrad State University)

SUBMITTED: November 20, 1957

Card 2/2



AUTHOR: Ogloblin, K. A.

SOV/79-28-12-18/41

TITLE: Reaction of Nitrosyl Chloride With Unsaturated Hydrocarbons  
(Vzaimodeystviye khloristogo nitrozila s nenasyshchennymi ugle-  
vodorodami)  
IV. Reaction With Butene-1 (IV. Reaktsiya s butenom-1)

PERIODICAL: Zhurnal obshchey khimii, 1958, Vol 28, pp 3245-3248 (USSR)

ABSTRACT: Earlier (Ref 1) the author had found that in the reaction of nitrosyl chloride with 2-methyl butene-3 the 4-nitro-3-chloro-2-methyl butane and 3,4-dichloro-2-methyl butane are formed as main products. To prove the direction typical of this reaction also with other monosubstituted ethylene hydrocarbons the reaction of nitrosyl chloride with butene-1 was carried out in ether solution in sealed ampoules at room temperature. The high pressure often caused explosions. For this reason, some less concentrated solutions of the reagents in ether and a smaller amount of the reaction mixture were taken (one third of the ampoule filling). Butene-1 in this reaction is less reactive than 2-methyl butene-3. 1-nitro-2-chloro butane and dichloride-1,2-dichloro butane turned out to be the main products of butene-1 with nitrosyl chloride, and isopropyl ethylene. 1-nitroso-2-chloro butane formed in the first stage

Card 1/2

OGIOBLIN, K.A.

Reaction of nitrosyl chloride with unsaturated hydrocarbons.  
Part 3: Reaction with isopropylethylene. Zhur. ob. khim. 27  
no.9:2541-2545 S '57. (MIRA 11:3)

1. Khimicheskiy institut Leningradskogo gosudarstvennogo universiteta.  
(Nitrosyl chloride) (Ethylene)

**OGLOBLIN, K.A.; MARKINA, G.V.**

Research in molecular chromatography. Part 3. Separation of mixtures of nitrophenols and nitrophenylnitramines. Zhur. ob.khim. 25 no.8:1616-1619 Ag '55. (MLBA 9:2)

1. Nauchno-issledovatel'skiy khimicheskiy institut Leningradskogo gosudarstvennogo universiteta.  
(Chromatographic analysis) (Phenol) (Amines)